

Molarity:

- a quantitative description of solution concentration.
- Abbreviated M

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

Problems: Show all work and circle your final answer.

1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required?

$$4.00 \text{ M} = \frac{\text{moles of solute}}{12.0 \text{ L}} \quad \text{moles of solute} = \boxed{48.0 \text{ mol}}$$

2. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M?

$$\begin{aligned} ? \text{ L} &= 250 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.25 \text{ L} \\ 0.150 \text{ M} &= \frac{\text{moles of solute}}{0.25 \text{ L}} \quad \text{moles of solute} = \boxed{0.038 \text{ mol}} \end{aligned}$$

3. What is the molarity of a solution of HNO_3 that contains 12.6 grams HNO_3 in 1.0 L of solution?

$$\begin{aligned} ? \text{ mol HNO}_3 &= 12.6 \text{ g HNO}_3 \times \frac{1 \text{ mol HNO}_3}{63.0 \text{ g HNO}_3} = 0.200 \text{ mol HNO}_3 \\ \text{M} &= \frac{0.200 \text{ mol HNO}_3}{1.0 \text{ L}} = \boxed{0.200 \text{ M}} \end{aligned}$$

4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?

$$\begin{aligned} 0.700 \text{ M} &= \frac{\text{moles of solute}}{0.250 \text{ L}} \quad \text{moles of solute} = 0.175 \text{ mol} \\ ? \text{ g KNO}_3 &= 0.175 \text{ mol KNO}_3 \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = \boxed{17.7 \text{ g KNO}_3} \end{aligned}$$

Molarity Review Worksheet 64

J Spring



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Table of Contents Molarity Review Worksheet 64

1. Understanding the eBook Molarity Review Worksheet 64
 - The Rise of Digital Reading Molarity Review Worksheet 64
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Review Worksheet 64
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Review Worksheet 64
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Review Worksheet 64
 - Personalized Recommendations
 - Molarity Review Worksheet 64 User Reviews and Ratings
 - Molarity Review Worksheet 64 and Bestseller Lists
5. Accessing Molarity Review Worksheet 64 Free and Paid eBooks
 - Molarity Review Worksheet 64 Public Domain eBooks
 - Molarity Review Worksheet 64 eBook Subscription Services
 - Molarity Review Worksheet 64 Budget-Friendly Options

6. Navigating Molarity Review Worksheet 64 eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Review Worksheet 64 Compatibility with Devices
 - Molarity Review Worksheet 64 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Review Worksheet 64
 - Highlighting and Note-Taking Molarity Review Worksheet 64
 - Interactive Elements Molarity Review Worksheet 64
8. Staying Engaged with Molarity Review Worksheet 64
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Review Worksheet 64
9. Balancing eBooks and Physical Books Molarity Review Worksheet 64
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Review Worksheet 64
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity Review Worksheet 64
 - Setting Reading Goals Molarity Review Worksheet 64
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Review Worksheet 64
 - Fact-Checking eBook Content of Molarity Review Worksheet 64
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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