

Molarity Practice Worksheet

Find the molarity (concentration) of the following solutions: Molarity = mole/Liters
Volume must be in liters! 1 liter = 1000 ml

- 1) 2.5 moles of sodium chloride is dissolved to make 0.050 liters of solution.
- 2) 3.0 moles of silver chloride is dissolved in enough water to make a 2.0 liter solution
- 3) 0.60 moles of potassium nitrate is dissolved in 300 milliliters of solution
- 4.) 4.2 moles of lithium fluoride are dissolved in 200 milliliters of solution.
- 5.) 18 grams of hydrochloric acid, HCl is dissolved to make 0.500 liters of solution.
- 6) 33 grams of Beryllium hydride, BeH_2 , are dissolved in 6.0 Liters of solution.
- 7) 734 grams of lithium sulfate, Li_2SO_4 are dissolved to make 2500 mL of solution.
- 8) 9.0×10^{23} molecules of CCl_4 dissolved to make 500 mL of solution.
- 6) I have two solutions. In the first solution, 1.0 moles of sodium chloride is dissolved to make 1.0 liters of solution. In the second one, 1.0 moles of

Molarity And Molality Practice Worksheet Answers

Mary Louise Turgeon



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