

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

# Modern Chemistry Homework 15 1 Answers

**Joacim Rocklöv**



## Modern Chemistry Homework 15 1 Answers:

**Teching Of Chemistry: Modern Methods** Anjali Khirwadkar, 2004      *Modern Coordination Chemistry* Jeff Leigh, 2007-10-31 Coordination chemistry as we know it today has been shaped by major figures from the past one of whom was Joseph Chatt Beginning with a description of Chatt s career presented by co workers contemporaries and students this fascinating book then goes on to show how many of today s leading practitioners in the field working in such diverse areas as phosphines hydrogen complexes transition metal complexes and nitrogen fixation have been influenced by Chatt The reader is then brought right up to date with the inclusion of some of the latest research on these topics all of which serves to underline Chatt s continuing legacy Intended as a permanent record of Chatt s life work and influence this book will be of interest to lecturers graduate students researchers and science historians      Chemistry Homework for OCR A for Double and Separate Awards Gareth Pritchard, 2001 Howework activites for OCR A Chemistry specifications      Fundamentals of Chemistry: A Modern Introduction Frank Brescia, 2012-12-02 Fundamentals of Chemistry Third Edition introduces the reader to the fundamentals of chemistry including the properties of gases atomic and molecular weights and the first and second laws of thermodynamics Chemical equations and chemical arithmetic are also discussed along with the structure of atoms chemical periodicity types of chemical bonds and condensed states of matter This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists Separation and purification are covered in the first chapter while the following chapters focus on atomic and molecular weights stoichiometry the structure of atoms and types of chemical bonds The molecular orbital MO theory of bonding galvanic cells and chemical thermodynamics are considered next Separate chapters are devoted to MO theory of covalent and metallic bonding orbital hybridization intermolecular forces acids and bases ionic equilibrium calculations and polymers and biochemicals This monograph is intended for chemistry students      **Modern Federal Practice Digest** ,1961      *Modern Magnetic Resonance* Graham A. Webb, 2007-05-26 Modern Magnetic Resonance provides a unique and comprehensive resource on up to date uses and applications of magnetic resonance techniques in the sciences including chemistry biology materials food medicine pharmaceuticals and marine sciences The widespread appeal of MMR methods for revealing information at the molecular and microscopic levels is noted and examples are provided from the chemical and other sciences Until now there has been no single publication that covers all the areas encompassed by Modern Magnetic Resonance by bringing together the various techniques and their applications in many scientific areas the internationally renowned Editors have created a resource of broad appeal to the scientific community The book includes High resolution solid and liquid state NMR Low resolution NMR Solution State NMR Magnetic Resonance Imaging Electron Spin Resonance Many applications taken from all of the chemical and related sciences      *NMR in Supramolecular Chemistry* Miquel Pons, 1999-03-31 NMR is better suited than any other experimental technique for the characterization of supramolecular

systems in solution The presentations included here can be broadly divided into three classes The first class illustrates the state of the art in the design of supramolecular systems and includes examples of different classes of supramolecular complexes catenanes rotaxanes hydrogen bonded rosettes tubes capsules dendrimers and metal containing hosts The second class comprises contributions to NMR methods that can be applied to address the main structural problems that arise in supramolecular chemistry The third class includes biological supramolecular systems studied by state of the art NMR techniques

**Secondary Teaching Methods** Kenneth T. Henson,1981 **School Life** ,1962 **Modern Persian, Elementary Level** Iago Gocheleishvili,2024-04-16 Modern Persian Elementary Level is an innovative Persian language textbook It is intended for university level learners and features material for two consecutive semesters of elementary Persian The textbook implements the most recent trends in language instruction including the basic tenets of flipped learning and communicative language teaching methodology with a student centric approach to language instruction Strengthened by its contemporary real world topics high frequency structures and vocabulary thematic presentation of material a plethora of engaging speaking activities in each chapter designated listening reading and writing sections and integration of cultural material this textbook is a straightforward and culturally engaging way to acquire functional proficiency in spoken and written Persian Complete with a companion website with over 200 audio and video presentations an answer key a searchable audio dictionary and a special appendix for instructors that features classroom activity materials for the entire year this textbook is an innovative and modern language learning resource The textbook also comes in an E book format to make language learning accessible on the go wherever you are

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heavier elements The text concludes with coverage of applications of NMR to the analysis of complex mixtures natural products carbohydrates and nucleic acids all areas of activity for researchers working at the chemistry life sciences interface The book s unique format provides some theoretical insight into the NMR technique used indicating the most common pulse sequences The book draws upon several NMR methods that are resurging or currently hot in the field and indicates the specific pulse sequence used by various spectrometer manufacturers for each technique It examines the analysis of complex mixtures a feature not found in most books on this topic

**Basic Chemistry Concepts and Exercises** John Kenkel, 2011-07-08 Chemistry can be a daunting subject for the uninitiated and all too often introductory textbooks do little to make students feel at ease with the complex subject matter Basic Chemistry Concepts and Exercises brings the wisdom of John Kenkel s more than 35 years of teaching experience to communicate the fundamentals of chemistry in a practical down to earth manner Using conversational language and logically assembled graphics the book concisely introduces each topic without overwhelming students with unnecessary detail Example problems and end of chapter questions emphasize repetition of concepts preparing students to become adept at the basics before they progress to an advanced general chemistry course Enhanced with visualization techniques such as the first chapter s mythical microscope the book clarifies challenging abstract ideas and stimulates curiosity into what can otherwise be an overwhelming topic Topics discussed in this reader friendly text include Properties and structure of matter Atoms molecules and compounds The Periodic Table Atomic weight formula weights and moles Gases and solutions Chemical equilibrium Acids bases and pH Organic chemicals The appendix contains answers to the homework exercises so students can check their work and receive instant feedback as to whether they have adequately grasped the concepts before moving on to the next section Designed to help students embrace chemistry not with trepidation but with confidence this solid preparatory text forms a firm foundation for more advanced chemistry training

**Chemistry** John Kenkel, Paul B. Kelter, David S. Hage, 2000-09-21 What a great idea an introductory chemistry text that connects students to the workplace of practicing chemists and chemical technicians Tying chemistry fundamentals to the reality of industrial life Chemistry An Industry Based Introduction with CD ROM covers all the basic principles of chemistry including formulas and names chemical bon

*The Publishers Weekly*, 1918

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