

Chapter Test B

Chapter: Chemical Equations and Reactions

Write a correct name (including oxidation state of the metal in ionic formulas) completed with reactant-product arrows and states.

- _____ 1. The product of a slightly soluble substance is represented as a double displacement reaction. Complete the equation of PbCl_2 .
a. precipitate
b. insoluble substance
c. soluble
- _____ 2. The balance chemical equation is represented as follows: the **a. products**
b. reactants
c. structure of the products
d. number of products
- _____ 3. In chemical equation, the aq refers to **a. water** **b. dissolved in water**
c. solid
d. insoluble
- _____ 4. The following list of substances are listed below as they are. Select the substance that is soluble in the following list: **a. sodium chloride**
b. potassium chloride
c. sodium chloride
d. sodium chloride
- _____ 5. The following list of chemical equations are listed below as they are. Select the equation that is balanced: **a. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$**
b. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2$
c. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2 + \text{O}_2$
d. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2 + \text{O}_2 + \text{O}_2$
- _____ 6. The following list of chemical equations are listed below as they are. Select the equation that is balanced: **a. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$**
b. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2$
c. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2 + \text{O}_2$
d. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2 + \text{O}_2 + \text{O}_2$
- _____ 7. The following list of chemical equations are listed below as they are. Select the equation that is balanced: **a. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$**
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d. $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2 + \text{O}_2 + \text{O}_2$

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