



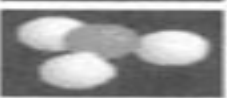










No. of Electron Pairs	Electron Pair Geometry (Bond Angle)	No. of Pendant Atoms	Molecular Geometry	Example Formula	Image
2	linear (180°)	2	linear	BeH ₂	
3	trigonal planar (120°)	3	trigonal planar	CO ₃ ²⁻	
		2	bent	NO ₂ ⁻	
4	tetrahedral (109.5°)	4	tetrahedral	CH ₄	
	(107 deg)	3	trigonal pyramidal	NH ₃	
	(104.5 deg)	2	bent	H ₂ O	
5	trigonal bipyramidal (90°, 120°)	5	trigonal bipyramidal	PCl ₅	
		4	see-saw	SF ₄	
		3	T-shaped	BrF ₃	
		2	linear	ICl ₂ ⁻	
6	octahedral (90°)	6	octahedral	SF ₆	
		5	square pyramidal	BrF ₅	
		4	square planar	ICl ₄ ⁻	

Modern Chemistry Chapter 8 1 Review Answers

Rachel S Tattersall



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