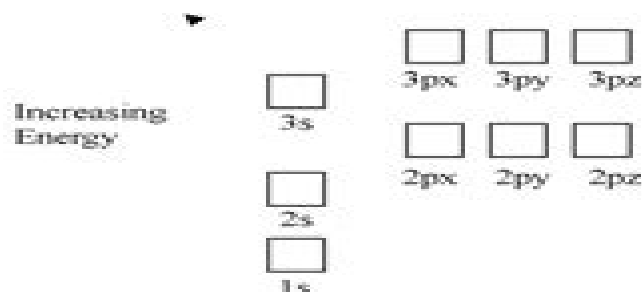


Electron configuration

The area in which there is 90% probability of finding an electron is called an orbital. An orbital can hold two electrons.

These orbitals can be represented in the following way:

Aufbau Diagram



Each box represents an orbital. Electrons are designated by arrows pointing up or down which represent clockwise or counterclockwise spins.



To determine where the electrons of a particular element are located you use the following rules:

RULE 1: An electron will always be found in the lowest energy orbital available that has room.

Therefore the electron configuration of Hydrogen is: This is designated as: $1s^1$

RULE 2: An orbital may hold a maximum of 2 electrons.

RULE 3: The Pauli Exclusion Principle – Two electrons sharing an orbital will have opposite spins.

Therefore the electron configuration of Helium is: This is designated as: $1s^2$

1. Draw an orbital diagram and write the electron configuration for Lithium, Beryllium, and Boron.



Electron Configuration Packet Key

L Darling-Hammond

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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