

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium *note: no d's, so the p. table referenced has Ac as 5f*
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^3$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

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Table of Contents Electron Configuration Worksheet And Lots More Answers

1. Understanding the eBook Electron Configuration Worksheet And Lots More Answers
 - The Rise of Digital Reading Electron Configuration Worksheet And Lots More Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Worksheet And Lots More Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Worksheet And Lots More Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Electron Configuration Worksheet And Lots More Answers
 - Personalized Recommendations
 - Electron Configuration Worksheet And Lots More Answers User Reviews and Ratings
 - Electron Configuration Worksheet And Lots More Answers and Bestseller Lists
5. Accessing Electron Configuration Worksheet And Lots More Answers Free and Paid eBooks
 - Electron Configuration Worksheet And Lots More Answers Public Domain eBooks
 - Electron Configuration Worksheet And Lots More Answers eBook Subscription Services
 - Electron Configuration Worksheet And Lots More Answers Budget-Friendly Options
6. Navigating Electron Configuration Worksheet And Lots More Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Worksheet And Lots More Answers Compatibility with Devices
 - Electron Configuration Worksheet And Lots More Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Worksheet And Lots More Answers
 - Highlighting and Note-Taking Electron Configuration Worksheet And Lots More Answers
 - Interactive Elements Electron Configuration Worksheet And Lots More Answers
8. Staying Engaged with Electron Configuration Worksheet And Lots More Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Worksheet And Lots More Answers
9. Balancing eBooks and Physical Books Electron Configuration Worksheet And Lots More Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Worksheet And Lots More Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Worksheet And Lots More Answers
 - Setting Reading Goals Electron Configuration Worksheet And Lots More Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Electron Configuration Worksheet And Lots More Answers
 - Fact-Checking eBook Content of Electron Configuration Worksheet And Lots More Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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