

Thermodynamics

MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question.

1. _____ If an ideal gas does positive work on its surroundings,
 - a. the temperature of the gas increases.
 - b. the volume of the gas increases.
 - c. the pressure of the gas increases.
 - d. the internal energy of the gas decreases.

2. _____ Sunlight strikes an ice cube at its melting point and causes it to melt. What is this process?
 - a. adiabatic process
 - b. isothermal process
 - c. isobaric process
 - d. isovolumetric process

3. _____ Which of the following processes for an ideal gas system has an unchanging internal energy and a heat intake that corresponds to the value of the work done by the system?
 - a. isothermal process
 - b. isobaric process
 - c. adiabatic process
 - d. isovolumetric process

4. _____ How is conservation of internal energy expressed for an isovolumetric system?
 - a. $Q = W = \Delta U = 0$; $U_i = U_f$
 - b. $Q = 0$; so $\Delta U = -W$
 - c. $\Delta T = 0$, so $\Delta U = Q - W = 0$, so $Q = W$
 - d. $\Delta V = 0$, so $P\Delta V = 0$ and $W = 0$, so $\Delta U = Q$

5. _____ A 4 mol ideal gas system undergoes an adiabatic process in which it expands and does 20 J of work on its environment. What is its change in internal energy?
 - a. -20 J
 - b. -5 J
 - c. 0 J
 - d. 20 J

6. _____ According to the second law of thermodynamics, the heat received by a heat engine operating in a complete cycle from a high-temperature reservoir
 - a. must be completely converted to work.
 - b. equals the entropy increase.
 - c. can be completely converted to internal energy.
 - d. cannot be completely converted to work.

Holt Physics Chapter 11 Test

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