

## Chemistry Unit 7 Worksheet 4

### Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedures outlined in class. Be sure to circle your final answer.

- Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$ , that must decompose to produce potassium chloride,  $\text{KCl}$ , and 1.0 moles of oxygen gas.
- In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 0.40 g of magnesium?
- Ethanol,  $\text{C}_2\text{H}_5\text{OH}$ , reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.00 moles of ethanol?
- Determine the mass of sodium chloride produced when 0.70 g of sodium (Na) reacts with sodium hydroxide according to the following equation:  
 $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- In the copper-silver nitrate lab copper metal and silver nitrate solution reacted to produce silver metal and copper(II) nitrate in solution.  
 A student placed a copper wire with a mass of 0.00 g in the reaction test tube. The silver-silver nitrate solution contained 0.40 g of silver nitrate.  
 She obtained 0.07 g of silver metal. Calculate the percent yield of silver.
- When hydrochloric acid (HCl) is added to sodium hydroxide solution, the products are water, aqueous sodium chloride and sodium chloride gas. What is the per cent yield of 0.00 g of  $\text{CCl}_4$  are collected when 0.00 g of sodium hydroxide solution reacts with excess HCl?
- Phosphorus and bromine react (chemically) together to form phosphorus tribromide. If 0.0 g of phosphorus and 0.0 g of bromine react, how many grams of  $\text{PBr}_3$  could be produced?
- Iron sulfate and oxygen gas react to form iron oxide and sulfur dioxide. Determine the amount of  $\text{SO}_2}$  that should be produced in a reaction between 0.0 g of  $\text{FeSO}_4$  and 0.0 g of oxygen. What is the percent of the experiment?

$$1. \text{ 1.0 moles of } \text{KClO}_3 \quad 2. \text{ 0.000 moles of } \text{H}_2\text{O} \quad 3. \text{ 0.000 g } \text{H}_2\text{O} \quad 4. \text{ 0.000 g of NaCl}$$

$$5. \text{ 0.00 g} \quad 6. \text{ 0.00 g} \quad 7. \text{ 0.00 g of } \text{CCl}_4 \quad 8. \text{ 0.00 g of } \text{SO}_2 \quad 9. \text{ 0.00 g of } \text{FeSO}_4$$

# Modeling Chemistry Unit 7 Ws 4 V

**G Thomas**



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## **Table of Contents Modeling Chemistry Unit 7 Ws 4 V**

1. Understanding the eBook Modeling Chemistry Unit 7 Ws 4 V
  - The Rise of Digital Reading Modeling Chemistry Unit 7 Ws 4 V
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 7 Ws 4 V
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 7 Ws 4 V
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 Ws 4 V
  - Personalized Recommendations
  - Modeling Chemistry Unit 7 Ws 4 V User Reviews and Ratings
  - Modeling Chemistry Unit 7 Ws 4 V and Bestseller Lists
5. Accessing Modeling Chemistry Unit 7 Ws 4 V Free and Paid eBooks
  - Modeling Chemistry Unit 7 Ws 4 V Public Domain eBooks

- Modeling Chemistry Unit 7 Ws 4 V eBook Subscription Services
- Modeling Chemistry Unit 7 Ws 4 V Budget-Friendly Options
- 6. Navigating Modeling Chemistry Unit 7 Ws 4 V eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry Unit 7 Ws 4 V Compatibility with Devices
  - Modeling Chemistry Unit 7 Ws 4 V Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 Ws 4 V
  - Highlighting and Note-Taking Modeling Chemistry Unit 7 Ws 4 V
  - Interactive Elements Modeling Chemistry Unit 7 Ws 4 V
- 8. Staying Engaged with Modeling Chemistry Unit 7 Ws 4 V
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry Unit 7 Ws 4 V
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 Ws 4 V
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry Unit 7 Ws 4 V
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 7 Ws 4 V
  - Setting Reading Goals Modeling Chemistry Unit 7 Ws 4 V
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 7 Ws 4 V
  - Fact-Checking eBook Content of Modeling Chemistry Unit 7 Ws 4 V
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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