

100) Carbon tetrachloride consists of one carbon and four chlorine atoms. How many chlorine atoms are present in 1.00 mole of  $\text{CCl}_4$  molecules?

$$\text{one mole of } \text{CCl}_4 \text{ contains } 4 \text{ moles of } \text{Cl} \text{ atoms}$$

$$1 \text{ mole } \text{CCl}_4 \times \frac{4 \text{ moles Cl}}{1 \text{ mole } \text{CCl}_4} = 4 \text{ moles Cl}$$

$$\text{answer} = 4 \text{ moles Cl} \quad \text{100}$$

101) There are 4 monatomic gases in each buffer gas used for Phosor.

102) Describe how each of these gases is used.

Helium - for cooling / insulation  
 Argon - for shielding / welding  
 Neon - for lighting / display  
 Krypton - for lighting / display

103) Describe how each of these gases is used in a gas mixture.

Helium - for cooling / insulation  
 Argon - for shielding / welding  
 Neon - for lighting / display  
 Krypton - for lighting / display

104) Carbon tetrachloride is a compound of carbon and chlorine. How many chlorine atoms are present in 1.00 mole of  $\text{CCl}_4$  molecules?

105) How many chlorine atoms are present in 1.00 mole of  $\text{CCl}_4$  molecules?

106) What is the empirical formula of carbon tetrachloride?

$$\text{CCl}_4$$



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**Sabine Zange**



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