

## **Like Dissolves Like Practice Answers**

Using your knowledge of the "like dissolves like" principle of solubility, determine which component will dissolve best in each of the following solvents. Explain your reason for choosing each like dissolves.

1) Solvent: methanol (CH<sub>3</sub>OH)

Components: carbon tetrachloride and oxygen dichloride

Both oxygen dichloride and methanol are polar, while carbon tetrachloride is not.

2) Solvent: carbon dioxide

Components: methanol (CH<sub>3</sub>OH) and nitrogen dioxide

While both components are polar, methanol is considerably more polar than nitrogen dioxide (both components always have that charge separation, while nitrogen dioxide, on most of time, the not).

3) Solvent: water

Components: carbon tetrachloride and ammonia

Both water and ammonia are polar. Carbon tetrachloride is not.

4) Solvent: hexagonal alcohol

Components: water and methanol

Both water and hexagonal alcohol are polar. Methanol is not.

# Like Dissolves Like Practice Answer Key

**Silberberg**



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