

Gravimetric Analysis of a Chloride Salt (G)

Date submitted: 10/12/2021

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Lab section: LW3

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Purpose:

The purpose of the lab is to find the amount of chloride present in any unknown soluble salt, by showing and explaining the techniques present in gravimetric analysis.

Theory:

- i. This is the equation that describes the reaction between silver ion and chloride ion to form silver chloride:



- ii. AgCl solubility in water and the K_{sp} :

"Silver chloride is very insoluble, so we can assume that the above reaction goes to completion."¹

$$K_{sp} = 1.6 \times 10^{-10}$$

- iii. AgCl is precipitated in an acidic medium to reduce the chance of acid ions interfering with the results. These ions will form co-precipitates with silver in acid. Due to these co-precipitates, the results would be higher.
- iv. In order to ensure the precipitate does not go through the filter paper, which would give us lower results. To prevent this from happening, it is heated to make coagulate it into a clumpy mass. Heating the precipitate makes it coagulate, making it much harder for any precipitate to go through the filter paper.
- v. Silver chloride decomposes into silver and chloride in the presence of light, which is described by the following equation:



If this photodecomposition takes place in air, it will result in lower results since chlorine will be released as a gas. If it takes place in the presence of excess silver ion in aqueous solution, it will result in higher results, due to chlorine reacting with the silver to form silver chloride and solid silver from the decomposition remaining in the precipitate, which will make the results high. The following chemical equation is a result of chlorine reacting with silver to form silver chloride:

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Ian Pickup



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Table of Contents Gravimetric Analysis Of Chloride Salt Lab Report

1. Understanding the eBook Gravimetric Analysis Of Chloride Salt Lab Report
 - The Rise of Digital Reading Gravimetric Analysis Of Chloride Salt Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Gravimetric Analysis Of Chloride Salt Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gravimetric Analysis Of Chloride Salt Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gravimetric Analysis Of Chloride Salt Lab Report
 - Personalized Recommendations
 - Gravimetric Analysis Of Chloride Salt Lab Report User Reviews and Ratings
 - Gravimetric Analysis Of Chloride Salt Lab Report and Bestseller Lists
5. Accessing Gravimetric Analysis Of Chloride Salt Lab Report Free and Paid eBooks

- Gravimetric Analysis Of Chloride Salt Lab Report Public Domain eBooks
 - Gravimetric Analysis Of Chloride Salt Lab Report eBook Subscription Services
 - Gravimetric Analysis Of Chloride Salt Lab Report Budget-Friendly Options
6. Navigating Gravimetric Analysis Of Chloride Salt Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Gravimetric Analysis Of Chloride Salt Lab Report Compatibility with Devices
 - Gravimetric Analysis Of Chloride Salt Lab Report Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gravimetric Analysis Of Chloride Salt Lab Report
 - Highlighting and Note-Taking Gravimetric Analysis Of Chloride Salt Lab Report
 - Interactive Elements Gravimetric Analysis Of Chloride Salt Lab Report
 8. Staying Engaged with Gravimetric Analysis Of Chloride Salt Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gravimetric Analysis Of Chloride Salt Lab Report
 9. Balancing eBooks and Physical Books Gravimetric Analysis Of Chloride Salt Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gravimetric Analysis Of Chloride Salt Lab Report
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gravimetric Analysis Of Chloride Salt Lab Report
 - Setting Reading Goals Gravimetric Analysis Of Chloride Salt Lab Report
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gravimetric Analysis Of Chloride Salt Lab Report
 - Fact-Checking eBook Content of Gravimetric Analysis Of Chloride Salt Lab Report
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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