

2.5	The new series/<i>Die nuwe reeks</i> is $16 + 44 + 72 + \dots + 1752$ $16 + 28(n - 1) = 1752$ $1736 = 28(n - 1)$ $62 = n - 1$ $n = 63$ OR/OF $2 + 9 + \underline{16} + 23 + 30 + 37 + \underline{44} + 51 + \dots + \underline{1752}$ T_3 is divisible by /is deelbaar deur 4 Then $T_7, T_{11}, T_{15}, \dots, T_{251}$ are divisible by 4, thus each 4th term is divisible by 4. Daarna is $T_7, T_{11}, T_{15}, \dots, T_{251}$ deelbaar deur 4, d.w.s. elke 4^{de} term is deelbaar deur 4. $\therefore$ number of terms divisible by 4 will be $= \frac{251 - 3}{4} + 1 = 63$ $\therefore$ aantal terme deelbaar deur 4 sal wees $= \frac{251 - 3}{4} + 1 = 63$ OR/OF Position of terms divisible by 4: $3 ; 7 ; 11 ; \dots ; 247 ; 251$ $T_n = 4n - 1 = 251$ $4n = 252$ $n = 63$	✓✓ generating new series divisible by 4/ <i>vorming van nuwe reeks deelbaar deur 4</i> ✓ $T_n = 1752$ ✓ 63 (4) ✓ T_3 is divisible by 4/ <i>is deelbaar deur 4</i> ✓ identifying terms divisible by 4/ <i>identifiseer terme deelbaar deur 4</i> ✓ reasoning/redenering ✓ 63 (4)
	Position of terms divisible by 4: $3 ; 7 ; 11 ; \dots ; 247 ; 251$ $T_n = 4n - 1 = 251$ $4n = 252$ $n = 63$	✓✓ generating sequence involving position of terms/<i>vorming van reeks i.t.v. posisie van terme</i> ✓ $T_n = 251$ ✓ 63 (4)

Grade Maths Memo 2014 November

Victor M. Corman



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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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