



Periodic Table of the Elements

Legend:

- metals
- metalloids
- nonmetals
- halogens
- noble gases
- transition metals
- lanthanide series
- actinide series

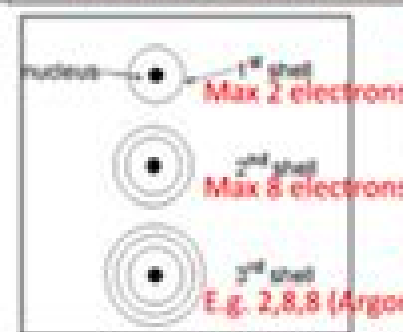
1	2											3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18							
H	He											Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar							
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn				
Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr	Rf	Db	Sg	Bh	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr	Rf	Db	Sg	Bh	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn

An atom:
 Number of Protons = Number of Electrons
 Mass number = Protons + Neutrons

C2 Chemistry

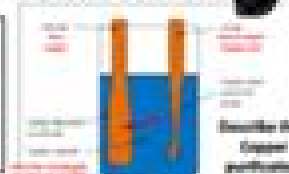
Periodic table arranges elements by
ATOMIC NUMBER (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
 E.g. Lithium 2,1 is a group 1 element.

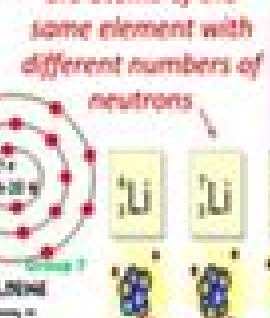
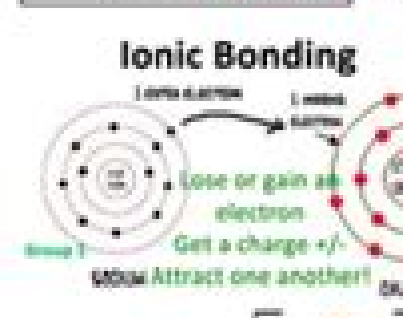
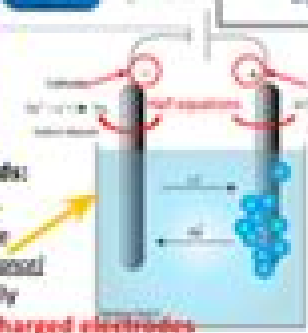


ISOTOPES
 are atoms of the same element with different numbers of neutrons

OIL RIG
 Oxidation is loss (of electrons)
 Reduction is gain (of electrons)



Electrolysis:
 Splitting up a substance using electricity

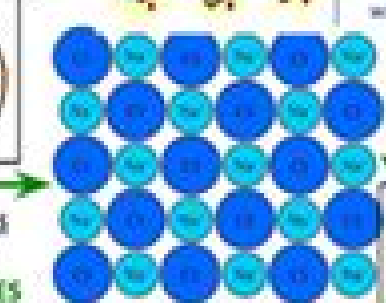


Metal ions only arranged in layers

ELECTROSTATIC ATTRACTION between +IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well



As a solid: High melting and Boiling Point
 Conducts electricity when melted or dissolved:
 Ions move freely and carry a charge

NANO-TECHNOLOGY
 Means **REALLY REALLY REALLY** small!!!!

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Philippa Gardom Hulme



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 - The Rise of Digital Reading Gcse Additional Chemistry C2 Revision Booklet
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Additional Chemistry C2 Revision Booklet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Additional Chemistry C2 Revision Booklet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Additional Chemistry C2 Revision Booklet
 - Personalized Recommendations
 - Gcse Additional Chemistry C2 Revision Booklet User Reviews and Ratings
 - Gcse Additional Chemistry C2 Revision Booklet and Bestseller Lists

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 - Gcse Additional Chemistry C2 Revision Booklet Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Gcse Additional Chemistry C2 Revision Booklet
 - Highlighting and Note-Taking Gcse Additional Chemistry C2 Revision Booklet
 - Interactive Elements Gcse Additional Chemistry C2 Revision Booklet
8. Staying Engaged with Gcse Additional Chemistry C2 Revision Booklet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Additional Chemistry C2 Revision Booklet
9. Balancing eBooks and Physical Books Gcse Additional Chemistry C2 Revision Booklet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Additional Chemistry C2 Revision Booklet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gcse Additional Chemistry C2 Revision Booklet
 - Setting Reading Goals Gcse Additional Chemistry C2 Revision Booklet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gcse Additional Chemistry C2 Revision Booklet
 - Fact-Checking eBook Content of Gcse Additional Chemistry C2 Revision Booklet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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