

# 5

## ELECTRONS IN ATOMS

### Practice Problems

In your notebook, solve the following problems.

#### SECTION 5.1 MODELS OF THE ATOM

- How many sublevels are in the following principal energy levels?
 

|            |            |            |
|------------|------------|------------|
| a. $n = 1$ | c. $n = 3$ | e. $n = 5$ |
| b. $n = 2$ | d. $n = 4$ | f. $n = 6$ |
- How many orbitals are in the following sublevels?
 

|                |                |                                 |
|----------------|----------------|---------------------------------|
| a. 1s sublevel | d. 4f sublevel | g. fifth principal energy level |
| b. 5s sublevel | e. 7s sublevel | h. 6d sublevel                  |
| c. 4d sublevel | f. 3p sublevel |                                 |
- What are the types of sublevels and number of orbitals in the following energy levels?
 

|            |            |            |
|------------|------------|------------|
| a. $n = 1$ | c. $n = 3$ | e. $n = 5$ |
| b. $n = 2$ | d. $n = 4$ |            |

#### SECTION 5.2 ELECTRON ARRANGEMENT IN ATOMS

- Write a complete electron configuration of each atom.
 

|              |            |            |
|--------------|------------|------------|
| a. hydrogen  | d. barium  | g. krypton |
| b. vanadium  | e. bromine | h. arsenic |
| c. magnesium | f. sulfur  | i. radon   |

#### SECTION 5.3 PHYSICS AND THE QUANTUM MECHANICAL MODEL

- What is the wavelength of the radiation whose frequency is  $5.00 \times 10^{15} \text{ s}^{-1}$ ? In what region of the electromagnetic spectrum is this radiation?
- An inexpensive laser that is available to the public emits light that has a wavelength of 670 nm. What are the color and frequency of the radiation?
- What is the energy of a photon whose frequency is  $2.22 \times 10^{14} \text{ s}^{-1}$ ?
- What is the frequency of a photon whose energy is  $6.00 \times 10^{-19} \text{ J}$ ?
- Arrange the following types of electromagnetic radiation in order of increasing frequency.
 

|                |                  |                |
|----------------|------------------|----------------|
| a. infrared    | c. visible light | e. microwaves  |
| b. cosmic rays | d. radio waves   | f. ultraviolet |
- Suppose that your favorite AM radio station broadcasts at a frequency of 1600 kHz. What is the wavelength in meters of the radiation from the station?

# Electrons In Atoms Worksheet Chapter 5 Answers

**Victor M. Corman**



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