

# Lab 5: Molecular Models



Table 1: Molecular Models

Example Atom (central atom is underlined>	Number of Bonds on Central Atom (multiple bonds count as 1)	Unshared electron pairs on central atom	Regions of Electron Density	Molecular Geometry	Structure
<u>C</u> O <sub>2</sub>	2	0	2	Linear	$O=C=O$
<u>Al</u> Cl <sub>3</sub>	3	0	3	Trigonal Planar	$\begin{array}{c} \text{Cl} \\   \\ \text{Cl}-\text{Al}-\text{Cl} \end{array}$
H <u>2</u> <u>O</u>	2	2	4	Bent	$\begin{array}{c} \cdot\cdot \\   \\ \text{H}-\text{O}- \\   \quad \cdot\cdot \\ \text{H} \end{array}$
<u>N</u> H <sub>3</sub>	3	1	4	Trigonal pyramidal	$\begin{array}{c} \cdot\cdot \\   \\ \text{H}-\text{N}-\text{H} \\   \\ \text{H} \end{array}$
<u>C</u> Cl <sub>4</sub>	4	0	4	Tetrahedral	$\begin{array}{c} \text{Cl} \\   \\ \text{Cl}-\text{C}-\text{Cl} \\   \\ \text{Cl} \end{array}$
<u>P</u> Cl <sub>5</sub>	5	0	5	Trigonal bipyramidal	$\begin{array}{c} \text{Cl} \\   \\ \text{Cl}-\text{P}-\text{Cl} \\   \\ \text{Cl} \end{array}$
<u>S</u> F <sub>6</sub>	6	0	6	Octahedral	$\begin{array}{c} \text{F} \\   \\ \text{F}-\text{S}-\text{F} \\   \\ \text{F} \end{array}$

# Models Of Molecular Compounds Lab 2prentice Hall Answers

**JG Myers**



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