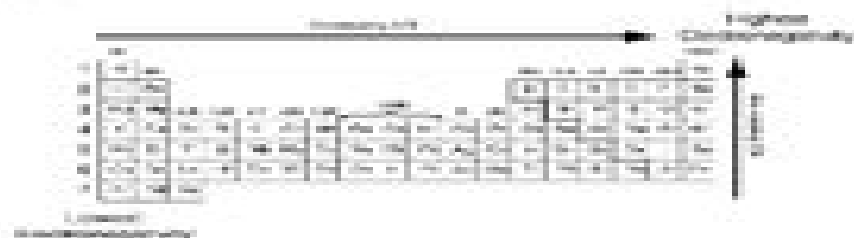


Unit 2 : Structure and Properties of Compounds

TYPES OF BONDING

- **Ionic Bonds:** one atom takes an electron from another atom
 - Between metal and nonmetal
 - Creates cations (+) and anions (-)
 - Atoms attracted to each other because of opposite charges
- **Covalent Bonds:** two atoms share electrons and have a strong attraction to electrons
 - Between nonmetal and nonmetal
 - Neither atom has a charge
 - Can be polar or non-polar
- **Metallic Bonds:** two atoms share electrons, but neither has a strong attraction to the electrons
 - Between metal and metal
 - Creates a "sea of electrons" between metals
- **Electronegativity:** the ability of an atom to attract an electron from another atom
 - Periodic Trend



- High electronegativity = requires a large amount of energy to remove an electron
- Low electronegativity = requires a small amount of energy to remove an electron

	Average EN	ΔEN
Ionic Bond	Medium	High
Metallic Bond	Low	Low
Covalent Bond	High	Low

LEWIS DIAGRAMS

- **Steps to creating a Lewis Diagram**
 - Determine the total number of valence electrons for all atoms (adjust for ions)
 - Determine the central atom, and place other atoms around joined with a single bond
 - Subtract electrons used for single bond from total
 - Add remaining electrons as dots to all the atoms in the molecule so that they have an octet (if possible)
 - If all valence electrons are assigned, but not all atoms have an octet, add double/triple bonds
 - If all atoms have an octet but not all electrons are used, extend the octet to the central atom
 - Atoms on the 3rd period and below can accommodate more than 8 electrons if needed (creates an extended octet)

Modeling Chemistry Unit 9 Review

Wenbin Ji



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