

II P.U.C PREPARATORY EXAM JAN/FEB - 2023

Time: 3 hours 15 minute

MATHEMATICS (35)

Max. Marks: 100

Instructions:

- 1) The question paper has five parts namely A, B, C, D and E. Answer all the parts.
- 2) Part A has 10 Multiple choice questions, 5 Fill in the blanks and 5 Very Short Answer questions of 1 mark each.
- 3) Part A should be answered continuously at one or two pages of Answer sheet and Only first answer is considered for the marks in subsection I and II of Part A.
- 4) Use the graph sheet for question on linear programming in PART E.

PART - A

I Answer ALL the Multiple Choice Questions

10 × 1 = 10

1. Let R be a relation in the set N given by $R = \{(a, b) : a - b = 2, b > 6\}$. Choose the correct answer
(A) $(2, 4) \in R$ (B) $(3, 8) \in R$ (C) $(6, 8) \in R$ (D) $(8, 7) \in R$
2. If $\sin^{-1} x = y$, then
(A) $0 \leq y \leq \pi$ (B) $-\frac{\pi}{2} \leq y \leq \frac{\pi}{2}$ (C) $0 < y < \pi$ (D) $-\frac{\pi}{2} < y < \frac{\pi}{2}$
3. The number of possible matrices of order 3×3 with each entry 0 or 1 is
(A) 27 (B) 18 (C) 81 (D) 512
4. Let A be a non singular square matrix of order 3×3 , then $|adj A|$ is equal to
(A) $|A|$ (B) $|A|^2$ (C) $|A|^3$ (D) $3|A|$
5. The function $f(x) = [x]$, where $[x]$ denotes the greatest integer function is continuous at
(A) 4 (B) -2 (C) 1 (D) 1.5
6. $\int \frac{\sin^2 x - \cos^2 x}{\sin^2 x \cos^2 x} dx$ is equal to
(A) $\tan x + \cot x + C$ (B) $\tan x - \cot x + C$ (C) $\tan x \cot x + C$ (D) $\tan x - \cot 2x + C$
7. If θ is the angle between any two vectors $\vec{a}$ and $\vec{b}$, then $|\vec{a} \cdot \vec{b}| = |\vec{a} \times \vec{b}|$ when θ is equal to
(A) 0 (B) $\frac{\pi}{2}$ (C) $\frac{\pi}{4}$ (D) π
8. A line makes equal angles with co-ordinate axes. Direction cosines of this line are
(A) $\pm(1, 1, 1)$ (B) $\pm(\frac{1}{\sqrt{3}}, \frac{1}{\sqrt{3}}, \frac{1}{\sqrt{3}})$ (C) $\pm(\frac{1}{3}, \frac{1}{3}, \frac{1}{3})$ (D) $\pm(\frac{1}{\sqrt{3}}, \frac{-1}{\sqrt{3}}, \frac{-1}{\sqrt{3}})$
9. Corner points of the feasible region determined by the system of constraints are $(0, 3)$, $(1, 1)$ and $(3, 0)$. Let $Z = px + qy$, where $p, q > 0$. Condition on p and q so that the minimum of Z occurs at $(3, 0)$ and $(1, 1)$ is
(A) $p = 2q$ (B) $p = \frac{q}{2}$ (C) $p = 3q$ (D) $p = q$
10. If $P(A) = 0.8$, $P(B) = 0.5$ and $P(B|A) = 0.4$, then $P(A \cap B)$ is equal to
(A) 0.12 (B) 0.20 (C) 0.32 (D) 0.9

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^-aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 $13\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$ because two NH_4^+aq

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