Chemistry Notes - Grade 10

5.1 Properties and Changes

Physical Change: Water to Ice Chemical Change: Hydrogen peroxide to Water

Physical Property:

Characteristic/description of the substance

4 Based on Qualitative Observations (senses)

4 Also based on Quantitative Observations (calculations and measurements) (e.g. colour, texture, density, smell)

Chemical Property:

Describes how a substance behaves in the presence of another substance (e.g. how it reacts with other elements and compounds)

Examples are flammability, corrosion, bleaching, rusting etc.

Physical Change:

No new substance created

4 Most can be reversed

6 E.g. Change of state, dissolving sugar in water, cutting paper

Chemical Change:

4 Produces a new substance

4 Most can't be reversed

4 E.g. Forest fire, boiling an egg, burning wood

Five Clues that a Chemical Change has Occurred:

1. A new colour is observed

A new odour is produced

3. Heat or light is produced or absorbed

4. Bubbles of gas are formed

5. A solid material (precipitate) forms in a liquid

5.4 Patterns in the Periodic Table

Where are the metals, non metals, and metalloids located on the periodic table?

4 Metals: Left side

4 Metalloid: On the staircase

4 Non Metals: Right side exception being Hydrogen

Where are the alkali metals found? What are their properties?

4 Location: Group 1

Properties: Highly Reactive, shiny, soft, low melting point, and they have one valence electrons, soft, ductile

Grade Chemistry Notes

Rachel Sandford

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