

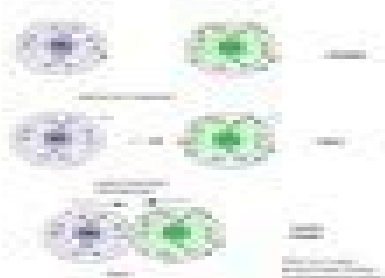
Final Study Guide

Valence Electrons: Electrons on the outermost shell or orbital, electrons involved in bonding, atoms are most stable when they have 8 (this is called the octet rule)

Isotopes: atoms with the same atomic number, but different atomic masses, because they have different amount of neutrons

Oxidation Numbers: positive or negative number shows the symbol that tells us how many electrons an atom has to lose or gain

Ionic Bond: a bond between a metal and a non-metal where electrons are lost and gained



Metal usually loses the electron and the nonmetal usually gains it, this causes the metal to be positively charged and the nonmetal to be negatively charged. Positives and negatives attract.

Properties:

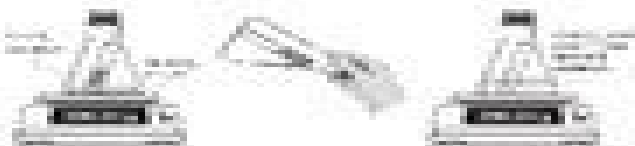
- Conduct electricity when melted or dissolved in water.
- High melting points
- Are polar
- Usually form crystalline structures

Covalent bonds: a bond between two nonmetals where electrons are shared, both nuclei are attracted to the shared electrons. Electrons can be shared, unequally which results in a polar covalent bond. Can also form double and triple bonds.

Properties:

- Low melting points
- Tend to be softer
- Usually don't dissolve in water
- Don't conduct electricity

The Law of Conservation of mass: Mass cannot be created or destroyed. The mass of the products must equal the mass of the reactants



as you can see in this picture the mass before the reaction is equal to the mass after the reaction.

This rule applies to atoms as well.

You must have the same number and type of atoms at the beginning of the reaction as you do after the reaction

Steps for Balancing Equations:

1. List the atoms involved in the reaction
2. Count how many of each type of atom on each side
3. Add coefficients to make the number of each type of atom on each side be equal
4. Recount the number of each type of atom on each side
5. If not equal repeat steps 3-4

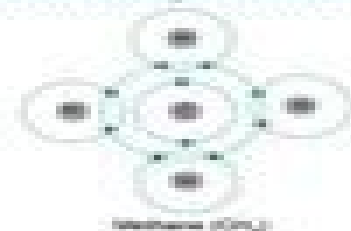
Structure of water molecule



Water

- Is a polar molecule

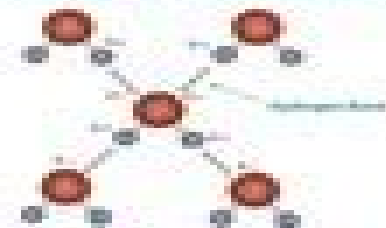
Electron sharing in covalent bond



Methane (CH₄)



	Reactants	Products
Hydrogen	4 H	4 H
Carbon	1 C	1 C
Oxygen	2 O	2 O
Chlorine	2 Cl	2 Cl
Compound	2 H ₂ O + CaCl ₂	CaCl ₂ + H ₂ + Cl ₂



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Table of Contents Ipcc Study Guide

1. Understanding the eBook Ipcc Study Guide
 - The Rise of Digital Reading Ipcc Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ipcc Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ipcc Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ipcc Study Guide
 - Personalized Recommendations
 - Ipcc Study Guide User Reviews and Ratings
 - Ipcc Study Guide and Bestseller Lists
5. Accessing Ipcc Study Guide Free and Paid eBooks
 - Ipcc Study Guide Public Domain eBooks
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 - Ipcc Study Guide Budget-Friendly Options

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 - IPC Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking IPC Study Guide
 - Interactive Elements IPC Study Guide
8. Staying Engaged with IPC Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers IPC Study Guide
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection IPC Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals IPC Study Guide
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of IPC Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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