#### Question 1 of 10

lonization energy increases down a group because the number of protons increases and it's more difficult to lose electrons because they are being pulled closer.

C False

C True

#### Question 2 of 10

Which of the following has the largest ionization energy?

CN

C Si

CAL

CP

#### Question 3 of 10

Which of the following elements has the lowest ionization energy?

Cesium (Cs)

C Sodium (Na)

C Lithium (Li)

C Rubidium (Rb)

#### Question 4 of 10

Order the elements in decreasing ionization energy. Na, S, Al, F, K, O

C Al, O, K, F, S, Na

F, O, S, Al, Na, K

Na, S, F, K, O, Al

# **Ionization Energy Exams Study Guide Answers**

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