

Question 1

(a) Using only the periodic table, arrange each set of atoms in order of increasing radius:

(1) Ba, Ca, Na



(2) In, Sn, As



(3) Cs, K, Al



(b) Which neutral atom is isoelectronic with each of the following ions?

(1) H^+

None

(The H^+ ion has 0 electrons. It is impossible for a neutral atom to have 0 electrons.)

(2) Zr^{4+}

Kr

(3) Mn^{7+}

Ar

(4) I^-

Xe

(c) Explain the following variations in atomic or ionic radii:

(1) $\text{I}^- > \text{I} > \text{I}^+$

As the atomic number stays constant and the number of electrons increases, the electron-electron repulsions increase. This leads to the electrons spreading out, and the anion becomes larger. The reverse is also true, so the cation is smaller.

(2) $\text{Ca}^{2+} > \text{Mg}^{2+} > \text{Be}^{2+}$

For cations with the same charge, ionic radii increase going down a column because of the increase in principal quantum number, which means the outer electrons are farther away from the nucleus on average. This leads to larger ion size.

Ionic Radius Answer Key

G Psacharopoulos

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents **Ionic Radius Answer Key**

1. Understanding the eBook **Ionic Radius Answer Key**
 - The Rise of Digital Reading **Ionic Radius Answer Key**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Ionic Radius Answer Key**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Ionic Radius Answer Key**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Ionic Radius Answer Key**
 - Personalized Recommendations
 - **Ionic Radius Answer Key** User Reviews and Ratings
 - **Ionic Radius Answer Key** and Bestseller Lists

5. Accessing Ionic Radius Answer Key Free and Paid eBooks
 - Ionic Radius Answer Key Public Domain eBooks
 - Ionic Radius Answer Key eBook Subscription Services
 - Ionic Radius Answer Key Budget-Friendly Options
6. Navigating Ionic Radius Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Radius Answer Key Compatibility with Devices
 - Ionic Radius Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Radius Answer Key
 - Highlighting and Note-Taking Ionic Radius Answer Key
 - Interactive Elements Ionic Radius Answer Key
8. Staying Engaged with Ionic Radius Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Radius Answer Key
9. Balancing eBooks and Physical Books Ionic Radius Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Radius Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Radius Answer Key
 - Setting Reading Goals Ionic Radius Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Radius Answer Key
 - Fact-Checking eBook Content of Ionic Radius Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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