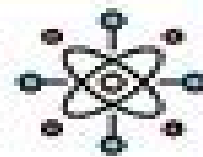


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules by sharing or transferring electrons. This process is called bonding. The resulting molecule is more stable than the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by a sea of delocalized electrons surrounding positive metal ions. Example: Fe , Cu , Al .

Properties of Bonds:

- Covalent bonds are strong and directional.
- Ionic bonds are strong but non-directional.
- Metallic bonds are strong and non-directional.

Science Real

A simplified periodic table of elements, color-coded by groups.

IONIC BONDING

Occurs between a metal and a non-metal. The metal atom loses one or more electrons to become a positively charged cation, and the non-metal atom gains those electrons to become a negatively charged anion. The resulting ionic bond is the electrostatic attraction between these oppositely charged ions.

COVALENT BONDING

Occurs between two non-metal atoms. They share one or more pairs of electrons to achieve a full outer shell. This shared pair of electrons is called a covalent bond.

Diagram illustrating Covalent Bonding:

A diagram showing two atoms, each with a nucleus and a shell of electrons. One electron from each atom is shared, forming a covalent bond. The resulting molecule has a full outer shell for both atoms.

Diagram illustrating Ionic Bonding:

A diagram showing a positively charged cation (yellow circle with a '+') and a negatively charged anion (red circle with a '-'). They are attracted to each other by an ionic bond.

Diagram illustrating Metallic Bonding:

A diagram showing a lattice of positive metal ions (yellow circles with '+') surrounded by a 'sea' of delocalized electrons (blue dots).

Properties of Bonds:

- Covalent bonds are strong and directional.
- Ionic bonds are strong but non-directional.
- Metallic bonds are strong and non-directional.

GUIDED READING ACTIVITY

Ionic And Metallic Bonding Guided Reading

Ying-Ying Zheng



Ionic And Metallic Bonding Guided Reading:

Ionic And Metallic Bonding Guided Reading Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Ionic And Metallic Bonding Guided Reading**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

https://staging.conocer.cide.edu/public/scholarship/fetch.php/Inner_Harbor.pdf

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