

A chemical bond - forms when 2 or more atoms rearrange valence
electrons to increase stability.

ionic bond - forms when valence electrons are transferred
from one atom to another

cation - atom loses electrons to become + charged

anion - atom gains electrons to become - charged

In ionic compounds the ions are arranged in a crystal lattice.
Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard - not easily crushed
- conduct electricity when melted or
dissolved because the ions are free to move.

covalent bond - electrons are shared, forming
molecules.

Covalent compounds have weaker forces holding the
molecules together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are volatile liquids or gases.
- soft - easier to crush
- are not conductors of electricity

electronegativity - property that tells how strong an atom's attraction
is for electrons.

Since oxygen has a higher electronegativity than hydrogen, oxygen
holds onto shared electrons tighter, giving the oxygen a partial
negative charge and the hydrogen a partial positive charge.

Ionic Bonds And Notetaking Guide Answer

D Siedentop



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Ionic Bonds And Notetaking Guide Answer Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Ionic Bonds And Notetaking Guide Answer**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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