

CHEMICAL BONDING

Answer the following questions.

The notation for sodium chloride, NaCl, stands for one

- ☒ (a) formula unit.
- ☐ (b) crystal.
- ☐ (c) atom.
- ☐ (d) molecule.

In a crystal of an ionic compound, each cation is surrounded by a number of

- ☐ (a) molecules.
- ☐ (b) positive ions.
- ☐ (c) dipoles.
- ☒ (d) negative ions.

Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is

- ☐ (a) higher in potential energy.
- ☒ (b) lower in potential energy.
- ☐ (c) equal in potential energy.
- ☐ (d) unstable.

The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?

- ☐ (a) Compound A is not an ionic compound.
- ☒ (b) It will be more difficult to break the bonds in compound A than those in compound B.
- ☐ (c) Compound B has larger crystals than compound A.
- ☐ (d) Compound A has larger crystals than compound B.

The forces of attraction between molecules in a molecular compound are generally

- ☐ (a) stronger than the attractive forces among formula units in ionic bonding.
- ☒ (b) weaker than the attractive forces among formula units in ionic bonding.
- ☐ (c) approximately equal to the attractive forces among formula units in ionic bonding.
- ☐ (d) equal to zero.

Describe the force that holds two ions together in an ionic bond.

The force of attraction between unlike charges holds a negative ion and a positive ion together in an ionic bond.

What type of energy best represents the strength of an ionic bond?

lattice energy

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Christian G. Meyer



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