

Name key

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. When an atom loses an electron it becomes a positive ion (cation), and when it gains an electron it becomes a negative ion (anion).
2. An element in the halogen group has 7 valence electrons, and wants to gain 1 more to have 8. This is because of the octet rule.
3. An ionic bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal.
4. In a covalent bond, atoms share electrons, whereas in an ionic bond, atoms transfer electrons.
5. (Honors only) When electrons are shared unequally in a covalent bond it is polar, and when they're shared equally it is nonpolar. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the electronegativities.
6. An atom can have up to 8 valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called noble gases.
7. Atoms can gain, lose, or share valence electrons to become more stable.

B. Formulas for Ionic compounds - Fill in the chart below

	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	<u>2</u>	<u>7</u>	<u>5</u>	<u>1</u>
OXIDATION #	<u>+2</u>	<u>-1</u>	<u>-3</u>	<u>+1</u>

2. What is the **chemical formula** for each of the following ionic compounds?
(*hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.*)

- a) Potassium and Chlorine - KCl
- b) Calcium and Bromine - CaBr₂
- c) Magnesium and oxygen - MgO

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PT Brinkman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Ionic Bonding Study Guide

1. Understanding the eBook Ionic Bonding Study Guide
 - The Rise of Digital Reading Ionic Bonding Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonding Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonding Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonding Study Guide
 - Personalized Recommendations
 - Ionic Bonding Study Guide User Reviews and Ratings
 - Ionic Bonding Study Guide and Bestseller Lists
5. Accessing Ionic Bonding Study Guide Free and Paid eBooks
 - Ionic Bonding Study Guide Public Domain eBooks
 - Ionic Bonding Study Guide eBook Subscription Services
 - Ionic Bonding Study Guide Budget-Friendly Options
6. Navigating Ionic Bonding Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonding Study Guide Compatibility with Devices
 - Ionic Bonding Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonding Study Guide
 - Highlighting and Note-Taking Ionic Bonding Study Guide
 - Interactive Elements Ionic Bonding Study Guide
8. Staying Engaged with Ionic Bonding Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Ionic Bonding Study Guide
- 9. Balancing eBooks and Physical Books Ionic Bonding Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonding Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Bonding Study Guide
 - Setting Reading Goals Ionic Bonding Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Bonding Study Guide
 - Fact-Checking eBook Content of Ionic Bonding Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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