

Ionic and Covalent Bonds Study Guide

Be able to define the following terms:

- | | | |
|-------------------|-----------------|--------------------|
| ➤ covalent bond | ➤ anion | ➤ polyatomic ion |
| ➤ Lewis structure | ➤ cation | ➤ hydrogen bonding |
| ➤ molecule | ➤ chemical bond | ➤ valence electron |
| ➤ polar bond | ➤ ionic bond | |
| ➤ non-polar bond | ➤ ion | |

1. A group of covalently bonded atoms that acts together as one charged atom is a polyatomic ion.
2. A chemical bond that occurs when atoms share electrons is a(n) covalent bond.
3. The octet rule tells you how many electrons an atom must gain, lose, or share to become stable.
4. The oxidation number of an atom is shown with a roman numeral in parenthesis.
5. How many electrons are needed in the outer energy levels of most atoms for the atom to be chemically stable? **8 (Eight)**
6. Why do the noble gases NOT form compounds readily? **They have 8 electrons in their outer electron shell.**
7. What is the total number of atoms in the compound $\text{Ca}(\text{ClO}_3)_2$? **9 (Nine)**
8. Explain why it is correct to say that a compound like calcium carbonate, CaCO_3 , contains both ionic and covalent bonds. **Carbonate is a polyatomic ion (covalent bonded) and then it forms an ionic bond with Calcium.**
9. Name or write the formulas for the following ionic compounds:
 - a. Al_2O_3 **Aluminum Oxide**
 - b. NH_4Cl **Ammonium Chloride**
 - c. FeO **Iron (II) Oxide**
 - d. manganese (V) oxide **Mn_2O_5**
 - e. calcium chloride **CaCl_2**
 - f. lead (IV) nitrate **$\text{Pb}(\text{NO}_3)_4$**

Ionic And Covalent Bonding Study Guide

Raffaella Di Napoli



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Table of Contents Ionic And Covalent Bonding Study Guide

1. Understanding the eBook Ionic And Covalent Bonding Study Guide
 - The Rise of Digital Reading Ionic And Covalent Bonding Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic And Covalent Bonding Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic And Covalent Bonding Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic And Covalent Bonding Study Guide
 - Personalized Recommendations
 - Ionic And Covalent Bonding Study Guide User Reviews and Ratings
 - Ionic And Covalent Bonding Study Guide and Bestseller Lists
5. Accessing Ionic And Covalent Bonding Study Guide Free and Paid eBooks

- Ionic And Covalent Bonding Study Guide Public Domain eBooks
- Ionic And Covalent Bonding Study Guide eBook Subscription Services
- Ionic And Covalent Bonding Study Guide Budget-Friendly Options
- 6. Navigating Ionic And Covalent Bonding Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic And Covalent Bonding Study Guide Compatibility with Devices
 - Ionic And Covalent Bonding Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic And Covalent Bonding Study Guide
 - Highlighting and Note-Taking Ionic And Covalent Bonding Study Guide
 - Interactive Elements Ionic And Covalent Bonding Study Guide
- 8. Staying Engaged with Ionic And Covalent Bonding Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic And Covalent Bonding Study Guide
- 9. Balancing eBooks and Physical Books Ionic And Covalent Bonding Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic And Covalent Bonding Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic And Covalent Bonding Study Guide
 - Setting Reading Goals Ionic And Covalent Bonding Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic And Covalent Bonding Study Guide
 - Fact-Checking eBook Content of Ionic And Covalent Bonding Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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