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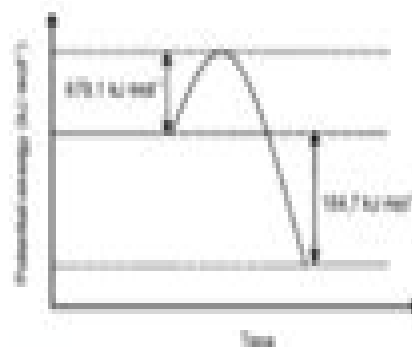
CHEMICAL CHANGE

11

ENDOTHERMIC & EXOTHERMIC REACTIONS

QUESTION 1 (start on a new page)

The diagram shows the potential energy changes during the following chemical reaction:



- 6.1 Define activation energy. (2)
- 6.2 Is the reaction $2\text{H}_2\text{g} + \text{O}_2\text{g} \rightarrow 2\text{H}_2\text{Og}$ endo- or exothermic? (1)
Give a reason for the answer. (2)
- 6.3 What is the total bond energy (D_{H} and D_{O}) of the reactants? (2)
Give a reason for the answer. (2)
- 6.4 Determine the energy released by the bond formation of the H_2O molecule. (2)
- 6.5 What effect will the addition of a catalyst have on the value $581.7 \text{ kJ mol}^{-1}$? (1)
Write down only INCREASE, DECREASE or NO EFFECT. (1)
Give a reason for the answer. (2)

Investigate Endothermic Reaction Grade 11

Shasha Hu



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