

Name: _____

Genetics Practice Problems (Remote Edition)



1. For each genotype below, indicate whether it is heterozygous (**He**) or homozygous (**Ho**)

AA **Ho** Ee **He** Bb **He** ff **Ho**
Pp **He** gg **Ho** Dd **He** HH **Ho**

2. For each of the **genotypes** below determine what **phenotypes** would be possible.

Purple flowers are dominant to white

PP **Purple**
Pp **Purple**
pp **White**

Brown eyes are dominant to blue

BB **Brown**
Bb **Brown**
bb **Blue**

3. For each **phenotype** below, list the **genotypes** (remember to use the letter of the dominant trait)

Straight hair is dominant to curly

SS **straight**
Ss **straight**
ss **curly**

Tail spikes are dominant to plain tails

SS **spikes**
Ss **spikes**
ss **plain**



4. Complete the Punnett squares for each of the crosses listed below.

Tall (T) plants are dominant to short (t)

Tt x tt

	<u>T</u>	<u>t</u>
<u>t</u>	Tt	tt
<u>t</u>	Tt	tt

What percentage of the offspring are tall? **50%** short? **50%**

Tt x Tt

	<u>T</u>	<u>t</u>
<u>T</u>	TT	Tt
<u>t</u>	Tt	tt

What percentage of the offspring are tall? **75%** short? **25%**



5. A **homozygous tall** plant is crossed with a **short** plant.

Genetics Practice Problems 7 Answers

Rod Powers



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 $\text{GaBr} + \text{NaBr} \rightarrow \text{SnO}_4 + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{S}_3 + \text{SO}_3$
 $3\text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr} + 3\text{SNO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$ using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn} + \text{CuCl}_2$
 how many moles of $\text{Ce} + \text{ZnCl}_2 + \text{ZnCl}_2$ will be produced from 23.0 g of $\text{Ce} + \text{Zn}$ assuming $\text{Ce} + \text{CuCl}_2 + \text{CuCl}_2$ is available in excess

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 $\text{Br}_2 + \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4 + \text{Br}_2$

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 answer $\text{Ce} + \text{SiO}_2 \rightarrow 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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