

## Experiment 9

### Electrochemistry I – Galvanic Cell

**Introduction:** Chemical reactions involving the transfer of electrons from one reactant to another are called *oxidation-reduction* reactions or *redox* reactions. In a redox reaction, two *half-reactions* occur; one reactant gives up electrons (undergoes oxidation) and another reactant gains electrons (undergoes reduction). A piece of zinc going into a solution as zinc ions, with each Zn atom giving up 2 electrons, is an example of an oxidation half-reaction.



The oxidation number of  $\text{Zn}(s)$  is 0 and the oxidation number of the  $\text{Zn}^{2+}$  is +2. Therefore, in this half-reaction, **the oxidation number increases**, which is another way of defining an oxidation. In contrast, the reverse reaction, in which  $\text{Zn}^{2+}$  ions gain 2 electrons to become Zn atoms, is an example of reduction.



In a reduction there is a **decrease (or reduction) in oxidation number**. Chemical equation representing half-reactions must be both mass and charge balanced. In the half-reactions above, there is one zinc on both sides of the equation. The charge is balanced because the 2+ charge on the zinc ion is balanced by two electrons,  $2e^{-}$ , giving zero net charge on both sides.

Another example of reduction is the formation of solid copper from copper ions in solution.



In this half-reaction the oxidation number of the aqueous copper is +2, which decreases to 0 for the solid copper, and again charge and mass are balanced. However, no half-reaction can occur by itself. A redox reaction results when an oxidation and a reduction half-reaction are combined to complete a transfer of electrons as in the following example:



The electrons are not shown because they are neither reactants nor products but have simply been transferred from one species to another (from Zn to  $\text{Cu}^{2+}$  in this case). In this redox reaction, the  $\text{Zn}(s)$  is referred to as the *reducing agent* because it causes the  $\text{Cu}^{2+}$  to be reduced to Cu. The  $\text{Cu}^{2+}$  is called the *oxidizing agent* because it causes the  $\text{Zn}(s)$  to be oxidized to  $\text{Zn}^{2+}$ .

Any half-reaction can be expressed as a reduction as illustrated in the case where equation (1) can be reversed to equation (2). A measure of the tendency for a reduction to occur is its **reduction potential,  $E$** , measured in units of volts. At standard conditions, 25 °C and concentrations of 1.0 M for the aqueous ions, the measured voltage of the reduction half-reaction is defined as the **standard reduction potential,  $E^{\circ}$** . Standard reduction potentials

# Galvanic Cell Experiment Report

**Yi-Tong Ma**



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## **Table of Contents Galvanic Cell Experiment Report**

1. Understanding the eBook Galvanic Cell Experiment Report
  - The Rise of Digital Reading Galvanic Cell Experiment Report
  - Advantages of eBooks Over Traditional Books
2. Identifying Galvanic Cell Experiment Report
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Galvanic Cell Experiment Report
  - User-Friendly Interface
4. Exploring eBook Recommendations from Galvanic Cell Experiment Report
  - Personalized Recommendations
  - Galvanic Cell Experiment Report User Reviews and Ratings
  - Galvanic Cell Experiment Report and Bestseller Lists
5. Accessing Galvanic Cell Experiment Report Free and Paid eBooks

- Galvanic Cell Experiment Report Public Domain eBooks
- Galvanic Cell Experiment Report eBook Subscription Services
- Galvanic Cell Experiment Report Budget-Friendly Options
- 6. Navigating Galvanic Cell Experiment Report eBook Formats
  - ePub, PDF, MOBI, and More
  - Galvanic Cell Experiment Report Compatibility with Devices
  - Galvanic Cell Experiment Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Galvanic Cell Experiment Report
  - Highlighting and Note-Taking Galvanic Cell Experiment Report
  - Interactive Elements Galvanic Cell Experiment Report
- 8. Staying Engaged with Galvanic Cell Experiment Report
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Galvanic Cell Experiment Report
- 9. Balancing eBooks and Physical Books Galvanic Cell Experiment Report
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Galvanic Cell Experiment Report
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Galvanic Cell Experiment Report
  - Setting Reading Goals Galvanic Cell Experiment Report
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Galvanic Cell Experiment Report
  - Fact-Checking eBook Content of Galvanic Cell Experiment Report
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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