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3rd Edition

Chemical Bonds

2-1 If the isotope ^{32}S has 16 protons and 16 neutrons, how many protons and how many neutrons will the isotope ^{33}S have?

- 2-2** A. If 0.5 mole of glucose weighs 90 g, what is the molecular weight of glucose?
 B. What is the concentration, in grams per liter (g/l), of a 0.25 M solution of glucose?
 C. How many molecules are there in 1 mole of glucose?

2-3 Which of the following elements is LEAST abundant in living organisms?

- (a) Sulfur
 (b) Carbon
 (c) Oxygen
 (d) Nitrogen
 (e) Hydrogen

2-4 Your friend learns about Avogadro's number and thinks it is so huge that there may not even be a mole of living cells on Earth. You have recently heard that there are about 50 trillion (50×10^{12}) human cells in each adult human body, so you bet your friend \$5 that there is more than a mole of cells on Earth. Once you learn that each human contains more bacterial cells (in the digestive system) than human cells, you are sure that you have won the bet. The human population is now more than 6 billion (6×10^9). What calculation can you show your friend to convince him you are right?

2-5 Atoms form covalent bonds with each other by

- (a) sharing protons.
 (b) sharing electrons.
 (c) transferring electrons from one atom to the other.
 (d) sharing neutrons.
 (e) attraction of positive and negative charges.

2-6 A carbon atom contains six protons and six neutrons.

- A. What are its atomic number and atomic weight?
 B. How many electrons does it have?
 C. What is its valence? How does this affect carbon's chemical behavior?
 D. Carbon with an atomic weight of 14 is radioactive. How does it differ in structure from

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Table of Contents Essential Cell Biology 3rd Edition Practice Questions

1. Understanding the eBook Essential Cell Biology 3rd Edition Practice Questions
 - The Rise of Digital Reading Essential Cell Biology 3rd Edition Practice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Essential Cell Biology 3rd Edition Practice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Essential Cell Biology 3rd Edition Practice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Essential Cell Biology 3rd Edition Practice Questions
 - Personalized Recommendations
 - Essential Cell Biology 3rd Edition Practice Questions User Reviews and Ratings
 - Essential Cell Biology 3rd Edition Practice Questions and Bestseller Lists
5. Accessing Essential Cell Biology 3rd Edition Practice Questions Free and Paid eBooks

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 - Essential Cell Biology 3rd Edition Practice Questions eBook Subscription Services
 - Essential Cell Biology 3rd Edition Practice Questions Budget-Friendly Options
6. Navigating Essential Cell Biology 3rd Edition Practice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Essential Cell Biology 3rd Edition Practice Questions Compatibility with Devices
 - Essential Cell Biology 3rd Edition Practice Questions Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Essential Cell Biology 3rd Edition Practice Questions
 - Highlighting and Note-Taking Essential Cell Biology 3rd Edition Practice Questions
 - Interactive Elements Essential Cell Biology 3rd Edition Practice Questions
 8. Staying Engaged with Essential Cell Biology 3rd Edition Practice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Essential Cell Biology 3rd Edition Practice Questions
 9. Balancing eBooks and Physical Books Essential Cell Biology 3rd Edition Practice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Essential Cell Biology 3rd Edition Practice Questions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Essential Cell Biology 3rd Edition Practice Questions
 - Setting Reading Goals Essential Cell Biology 3rd Edition Practice Questions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Essential Cell Biology 3rd Edition Practice Questions
 - Fact-Checking eBook Content of Essential Cell Biology 3rd Edition Practice Questions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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