

Thus, the oxidation number of sodium in the compound is $1+$ because one electron was transferred from the sodium atom. Because an electron is transferred to the chlorine atom, its oxidation number is $1-$.

- 7- In the chemical formula for any ionic compound, the symbol of the cation is always written first, followed by the symbol of the anion.
- 8- Subscripts, which are small numbers to the lower right of a symbol, represent the number of ions of each element in an ionic compound. If no subscript is written, it is assumed to be one.

Write formulas for the ionic compounds formed by the following ions.

19. potassium and iodide 21. aluminum and bromide
20. magnesium and chloride 22. sodium and nitride

23. **Challenge:** Write the general formula for the ionic compound formed by elements from the two groups shown on the periodic table at the right.

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Formulas for polyatomic ionic compounds:

- 9- polyatomic ions are ions made up of more than one atom. A polyatomic ion acts as an individual ion. Thus, the formula for a polyatomic compound follows the same rules used for a binary compound. Because a polyatomic ion exists as a unit, never change subscripts of the atoms within the ion.

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