

Atoms and Bonding → Review and Reinforce

Ionic Bonds**Understanding Main Ideas**

Answer the following questions on a separate sheet of paper.

- How does an atom become a positive ion? How does an atom become a negative ion?
- How do ions form electrically neutral compounds?
- What characteristics do ionic compounds show?
- Why does the electrical conductivity of ionic compounds change when they are dissolved in water?

| Ions and Their Charges | | |
|------------------------|--------|--------------------|
| Name | Charge | Symbol/Formula |
| Ammonium | 1+ | NH_4^+ |
| Potassium | 1+ | K^+ |
| Calcium | 2+ | Ca^{2+} |
| Magnesium | 2+ | Mg^{2+} |
| Chloride | 1- | Cl^- |
| Oxide | 2- | O^{2-} |
| Sulfide | 2- | S^{2-} |
| Phosphate | 3- | PO_4^{3-} |

Use the chart above to answer the following on a separate sheet of paper.

- How many potassium ions are needed to balance the charge of one sulfide ion? Explain.
- Predict the formulas for calcium chloride and potassium phosphate.
- Name the following compounds: MgS , NH_4Cl , and K_2O .
- Which ions in the table are polyatomic ions?

Building Vocabulary

Answer the following questions on a separate sheet of paper.

- What is an ion?
- What is an ionic bond?
- What is the arrangement of ions in a crystal?



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G. Polaski, J. Means



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