B.Loursey dec		

Student Exploration: Ionic Bonds

Date:

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: chemical family, electron affinity, ion, ionic bond, metal, nonmetal, octet rule, shell, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

 Nate and Clara are drawing pictures with markers. There are 8 markers in a set. Nate has 9 markers and Clara has 7. What can Nate and Clara do so that each of them has a full set?

Give one marker to Clara

Maggie is sitting at a table with Fred and Florence. Maggie has 10 markers, but Fred and Florence each have only 7 markers. How can they share markers so each has 8?

Maggie can give 1 marker to Fred, and one marker to Florence.

Gizmo Warm-up

Just like students sharing markers, atoms sometimes share or swap electrons. By doing this, atoms form bonds. The *lonic Bonds* Gizmo allows you to explore how **ionic bonds** form.

To begin, check that Sodium (Na) and Chlorine (CI) are selected from the menus at right. Click Play (E) to see electrons orbiting the nucleus of each atom. (Note: These atom models are simplified and not meant to be realistic.)



Each atom consists of a central nucleus and several shells that contain electrons. The outermost electrons
are called valence electrons. (Inner electrons are not shown.)

How many valence electrons does each atom have? Sodium:

1

Chlorine: 7

- Click Pause (!L.). Elements can be classified as metals and nonmetals. Metals do not hold on to their valence electrons very tightly, while nonmetals hold their electrons tightly. Electron affinity is a measure of how tightly the valence electrons are held.
 - A. Try pulling an electron away from each atom. Based on this experiment, which atom is a metal?

Sodium

Which is a nonmetal?

Chlorine

B. Try moving an electron from the metal to the nonmetal. What happens?

Sodium's electron is transferred to Chlorine's outer ring of electrons.

Ionic Bond Gizmo Answer Key

Christian G. Meyer

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