

- Directions:**  
 1. An element may be assigned a positive or negative charge depending on whether it is a metal or nonmetal.  
 2. An atom with a positive charge is a cation.  
 3. An atom with a negative charge is an anion.  
 4. The number of protons in an atom is the atomic number.

# Ion Practice Set

1. Write the symbol for the ion that has the same number of protons as the element but a different number of electrons.  
 2. Write the symbol for the ion that has the same number of protons as the element but a different number of electrons.

Complete the following table, using the periodic table in the back of your book.

	Element Name	Ion Symbol	Number of Protons	Number of Electrons	Number of Electrons Lost or Gained	Charge
1	Sodium	$\text{Na}^+$	11	10	lost one	positive
2	Chlorine	$\text{Cl}^-$	17	18	gained one	negative
3	Calcium	$\text{Ca}^{2+}$	20	18	lost two	positive
4	Oxygen	$\text{O}^{2-}$	8	10	gained two	negative
5	Aluminum	$\text{Al}^{3+}$	13	10	lost three	positive
6	Sulfur	$\text{S}^{2-}$	16	18	gained two	negative
7	Hydrogen	$\text{H}^+$	1	0	lost one	positive
8	Lithium	$\text{Li}^+$	3	2	lost one	positive
9	Fluorine	$\text{F}^-$	9	10	gained one	negative
10	Magnesium	$\text{Mg}^{2+}$	12	10	lost two	positive
11	Strontium	$\text{Sr}^{2+}$	38	36	lost two	positive
12	Barium	$\text{Ba}^{2+}$	56	54	lost two	positive
13	Strontium	$\text{Sr}^{2+}$	38	36	lost two	positive
14	Calcium	$\text{Ca}^{2+}$	20	18	lost two	positive
15	Sodium	$\text{Na}^+$	11	10	lost one	positive
16	Chlorine	$\text{Cl}^-$	17	18	gained one	negative

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