

Chapter 2: Atoms, Molecules and Ions

Dalton's Atomic Theory (1805)

1. All matter consists of **atoms**; tiny indivisible particles of an element that cannot be created or destroyed.
2. Atoms of one element **cannot** be subdivided, created, or destroyed during chemical reactions. (wrong)
3. Atoms of an element are **identical** in mass and other properties and are different from the atoms of any other element. (wrong)
4. Compounds result from the chemical combination of a **specific ratio** of atoms of different elements.

Atomic Structure: Electrons

- J.J. Thomson (1897) Cathode Ray Tube Experiment
- Observations:
 - Ray bends in magnetic field
 - Ray bends toward positive plate in electric field
 - Ray is identical for any cathode
- Conclusions:
 - Ray consists of charged particles
 - Ray consists of negative particles
 - These particles are found in ALL matter

Cathode Ray Tube experiment using magnets led to the discovery of electrons

Thomson's Model of the Atom

- Plum-pudding model:
 - E- distributed throughout diffuse, positively charged "squishy" sphere



Atomic Structure: The Nucleus

- Rutherford's Experiment
 - Bombard a thin gold foil with positively charged particles to test Thomson's model of the atom
 - Observations:
 - Atom is made up of smaller subatomic particles
 - Protons
 - Neutrons
 - Electrons (mass of only ab 1/2000th of other particles)
 - Conclusions

General Chemistry Study Guide

Clemens Wendtner



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