

Chemistry Laboratory 11B - The Iodine-Clock Reaction Shawn Yoon

Purpose:

The purposes of this laboratory were to:

- observe and record the effect of changing the concentration of a reactant on the rate of reactions, and to
- observe and record the effect of changing the temperature of a system on the rate of a reaction.

Materials and Methods:

The materials and methods for this laboratory followed the procedures laid out in Experiment 11B on pages 149-170 of *Essential Experiments for Chemistry* by Morrison and Boyd with the following exceptions:

- only two beakers were used to make water baths instead of four,
- volumetric pipettes were used instead of dropping pipettes,
- 25 mm x 100 mm test tubes were used instead of 15 mm x 175 mm ones,
- four test tubes were used instead of eight, and
- solutions A and B were mixed by shaking the test tube slightly instead of pouring between both test tubes.

Observations:

Part I:

First, approximately 50 mL of solution A and 50 mL of solution B were obtained from the front of the laboratory in 100 mL beakers. According to the instructor, solution A contained 0.002M KIO_3 and solution B contained 0.0005M NaHSO_3 , H_2SO_4 , and starch. Solution A appeared very clear, its appearance no different than water while solution B appeared very cloudy, due to the starch in solution.

Then, four 25 mm x 100 mm test tubes were obtained from the laboratory and placed in a test tube rack. Using a volumetric pipette, 10 mL of solution A was transferred into a graduated

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Ian Pickup



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Table of Contents Essential Experiments For Chemistry Answer Key 12c

1. Understanding the eBook Essential Experiments For Chemistry Answer Key 12c
 - The Rise of Digital Reading Essential Experiments For Chemistry Answer Key 12c
 - Advantages of eBooks Over Traditional Books
2. Identifying Essential Experiments For Chemistry Answer Key 12c
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Essential Experiments For Chemistry Answer Key 12c
 - User-Friendly Interface
4. Exploring eBook Recommendations from Essential Experiments For Chemistry Answer Key 12c
 - Personalized Recommendations

- Essential Experiments For Chemistry Answer Key 12c User Reviews and Ratings
- Essential Experiments For Chemistry Answer Key 12c and Bestseller Lists
- 5. Accessing Essential Experiments For Chemistry Answer Key 12c Free and Paid eBooks
 - Essential Experiments For Chemistry Answer Key 12c Public Domain eBooks
 - Essential Experiments For Chemistry Answer Key 12c eBook Subscription Services
 - Essential Experiments For Chemistry Answer Key 12c Budget-Friendly Options
- 6. Navigating Essential Experiments For Chemistry Answer Key 12c eBook Formats
 - ePub, PDF, MOBI, and More
 - Essential Experiments For Chemistry Answer Key 12c Compatibility with Devices
 - Essential Experiments For Chemistry Answer Key 12c Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Essential Experiments For Chemistry Answer Key 12c
 - Highlighting and Note-Taking Essential Experiments For Chemistry Answer Key 12c
 - Interactive Elements Essential Experiments For Chemistry Answer Key 12c
- 8. Staying Engaged with Essential Experiments For Chemistry Answer Key 12c
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Essential Experiments For Chemistry Answer Key 12c
- 9. Balancing eBooks and Physical Books Essential Experiments For Chemistry Answer Key 12c
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Essential Experiments For Chemistry Answer Key 12c
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Essential Experiments For Chemistry Answer Key 12c
 - Setting Reading Goals Essential Experiments For Chemistry Answer Key 12c
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Essential Experiments For Chemistry Answer Key 12c
 - Fact-Checking eBook Content of Essential Experiments For Chemistry Answer Key 12c

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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