Multiple Choice Questions

- K₁ and K₂ are the velocity constants of forward and backward reactions.
 The equilibrium constant k of the reaction is
 - (A) $K_j \times K_j$

(B) $K_1 - K_2$

(C) $\frac{K_I}{K_I}$

 $(D) \frac{K_I + K_I}{K_I - K_I}$

- Sed. (C).
- Which of the following reactions will be favoured at low pressure?
 - (A) N, +0, -2NO

- (B) H₂ + I₂ → 2HI
- (C) $PCI_A \longrightarrow 2PCI_A + CI_A$
- (D) $N_j + 3H_j = 5NH_j$

- Sed. (C).
- For the reaction, A+B--+C+D, if active mass of A is kept constant and active mass of B is tripled, the rate of reaction will become
 - (A) three times

(B) six times

(C) eight times

(D) nine times

- Sol. (D).
- 4. The rate of $\frac{K_p}{K_s}$ for the reaction:

$$CO(g) + \frac{1}{2}O_{1}(g) = CO_{1}(g)$$
 is

(A), I,

(B) RT

 $(C) (RT)^{I/I}$

(D) $(RT)^{-\epsilon}$

- Sol. (D).
- 5. Consider the reaction,

in a closed container at equilibrium. What would be the effect of addition of CaCO, on the equilibrium concentration of CO2?

(A) Increases

tBY Unpredictable

(C) Decreases:

(D) Remains unaffected

- Sed. (106.)
- For the reaction, PG_g(g), PG_g(g) + G_g(g), the forward reaction at constant temperature is favoured by
 - (A) introducing an mert gas at constant volume
 - (B) introducing PCI(g) at constant volume
 - (C) introducing PCl+(g) at constant volume
 - (D) introducing Cl3(g) at constant volume
- Sol. (C).
- A liquid is in equilibrium with its vapour at its boiling point. On an average, the molecules in the two phases have equal
 - (A) intermolecules forces
 - (B) potential energy
 - (C) kimeric emergy
 - (D) both kinetic energy and potential energy
- Soul. (C).
- 8. The unit of equilibrium constant for the reversible reaction

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Sabine Zange

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