

CHEMISTRY 10TH - Multiple Choice Questions

→ CHEMICAL EQUILIBRIUM

CHAPTER# 01

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01. Which one of the following statements is false about dynamic equilibrium?
- (a) It takes place in a closed container
 - (b) Concentration of reactant and products are not changed
 - (c) Rate of the forward reaction is equal to the rate of the reverse reaction
 - (d) ✓ Equilibrium cannot be disturbed by any external stress
02. Consider the following reaction and indicate which of the following best describes equilibrium constant expression K_c . $4\text{NH}_3 + 5\text{O}_2 \rightleftharpoons 4\text{NO} + 6\text{H}_2\text{O}$
- (a) ✓ $K_c = \frac{[\text{NO}]^4 [\text{H}_2\text{O}]^6}{[\text{NH}_3]^4 [\text{O}_2]^5}$
 - (b) $K_c = \frac{[\text{NH}_3]^4 [\text{O}_2]^5}{[\text{NO}]^4 [\text{H}_2\text{O}]^6}$
 - (c) $K_c = \frac{[\text{NH}_3][\text{O}_2]}{[\text{H}_2\text{O}][\text{NO}]}$
 - (d) $K_c = \frac{[4\text{NO}][6\text{H}_2\text{O}]}{[4\text{NH}_3][5\text{O}_2]}$
03. A reaction which is never goes to completion is known as reversible reaction. Reversible reaction is represented by:
- (a) dotted lines
 - (b) single arrow
 - (c) ✓ double arrow
 - (d) double straight line
04. When the magnitude of K_c is small, indicates:
- (a) ✓ Reaction mixture contain most of the reactant
 - (b) Reaction mixture contain most of the product
 - (c) Reaction mixture contain almost equal amount of reactant and product
 - (d) Reaction goes to completion
05. For which system does the equilibrium constant, K_c has units of concentration?
- (a) ✓ $\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$
 - (b) $\text{N}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)}$
 - (c) $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$
 - (d) $\text{CO}_{2(g)} + \text{H}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_2\text{O}_{(l)}$

Multiple Choice Questions And Answers In Chemical Equilibrium

Judd E. Hollander



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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Multiple Choice Questions And Answers In Chemical Equilibrium**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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