

2. Magnesium carbonate reacts with hydrochloric acid.



- (a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce  $240\text{cm}^3$  of carbon dioxide, measured at room temperature and pressure?

Data: 1 mol of any gas occupies  $24.0\text{dm}^3$  at room temperature and pressure

Molar mass of magnesium carbonate =  $84.3\text{g mol}^{-1}$

- ☒ A 0.843 g  
☐ B 8.43 g  
☐ C 84.3 g  
☐ D 843 g

$$\begin{aligned}\text{Volume of CO}_2 &= 240\text{cm}^3 = 0.24\text{dm}^3 \\ \text{moles of CO}_2 &= \frac{\text{Volume in dm}^3}{24} = \frac{0.24}{24} = 0.01 \\ \text{MgCO}_3 &\text{ contains same moles as CO}_2 \\ \text{moles} &= \frac{\text{mass}}{\text{Mr}} ; \text{ mass} = \text{moles} \times \text{Mr} \\ &= 0.01 \times 84.3 \\ &= 0.843\text{g}\end{aligned}$$

- (b) What is the minimum mass of magnesium carbonate needed to neutralise  $50.0\text{cm}^3$  of  $0.250\text{mol dm}^{-3}$  hydrochloric acid?

- ☐ A 0.423 g  
☒ B 0.527 g  
☐ C 1.05 g  
☐ D 2.11 g

$$\begin{aligned}\text{Concentration of HCl} &= \frac{\text{moles}}{\text{Vol in dm}^3} \\ \text{moles} &= \text{Conc} \times \text{Vol in dm}^3 \\ &= 0.250 \times 0.05 \\ &= 0.0125\text{ moles of HCl}\end{aligned}$$

$$\begin{aligned}\text{MgCO}_3 : \text{HCl} \\ 1 : 2 \\ \frac{1}{2} : 1 \\ \frac{1}{2} \times 0.0125 : 0.0125 \\ 6.25 \times 10^{-3} : 0.0125 \\ \downarrow \qquad \qquad \downarrow \\ \text{MgCO}_3 \qquad \text{HCl} \\ \text{moles} \qquad \text{moles}\end{aligned}$$

- (c) What would be seen at the end of the reaction with excess acid?

- ☒ A A colourless solution  
☐ B A coloured solution  
☐ C A white precipitate  
☐ D A coloured precipitate

all  $\text{MgCO}_3$  is utilized  
and color of  
 $\text{MgCO}_2$  colorless

$$\text{mass} = \text{moles} \times \text{Mr}$$

(Total for Question 2 = 3 marks)

$$\begin{aligned}&= 6.25 \times 10^{-3} \times 84.3 \\ &= 0.527\text{g}\end{aligned}$$

# Multiple Choice Questions Of Stoichiometry

**M Mark**



## **Multiple Choice Questions Of Stoichiometry:**

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