

POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

	Metal or Nonmetal	Number of protons in the atom	Number of protons in the ion	Number of neutrons in the atom	Number of neutrons in the ion	Number of electrons in the atom	Number of electrons in the ion	Charge on the ion
Lithium	metal	3	3	4	4	3	2	1+
Magnesium	<i>Metal</i>	12	12	12	12	12	10	2+
Aluminum	<i>Metalloid</i>	13	13	14	14	13	10	3+
Fluorine	<i>Non metal</i>	9	9	10	10	9	10	1-
Oxygen	nonmetal	8	8	8	8	8	10	2-
Nitrogen	<i>Non metal</i>	7	7	7	7	7	10	3-

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

No

4. Which subatomic particle carries a positive charge?

Protons are positive

5. Which subatomic particle carries a negative charge?

Electrons are negative

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

The charge of an ion = # of protons - # of electrons

Ions Pogil Answers

Tomas Baer, Cheuk-Yiu Ng, Ivan Powis

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Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more **Reactions of Ions** William Edward Marsden, 1996 *Organic Chemistry* Suzanne M. Ruder, The POGIL Project, 2015-12-29 ORGANIC CHEMISTRY *Ions in Solution* Ronald Wilfrid Gurney, 1962 *Ions in Solution* R. W. Gurney, 1936 Ions in Solution Ronald Wilfrid Gurney, 1936 **Ions and Light** Michael T. Bowers, 2013-10-22 Gas Phase Ion Chemistry Volume 3 Ions and Light discusses how ions are formed by electron impact ion molecule reactions or electrical discharge This book discusses the use of light emitted by excited molecules to characterize either the chemistry that formed the excited ion the structure of the excited ion or both Organized into 10 chapters this volume begins with an overview of the extension of the classical flowing afterglow technique to include infrared and chemiluminescence and laser induced fluorescence detection This text then examines the experiments involving molecules that are isolated from collisions for periods exceeding several milliseconds Other chapters consider the photodetachment in negative ion beams and the chemical information that can be obtained from such studies This book discusses as well the electronic states of the open shell organic cations The final chapter deals with ion beam spectroscopy This book is a valuable resource for chemists and scientists **The Theory of Ions** William Tibbles, 1909 **Ions in Solution (1) (sound Recording)**. John Burgess, 1984 Gaseous Molecular Ions Eugen Illenberger, Jacques Momigny, 1992 Most of the matter in our solar system and probably within the whole universe exists in the form of ionized particles On the other hand in our natural environment gaseous matter generally consists of neutral atoms and molecules Only under certain conditions such as within the path of lightning or in several technical devices e.g. gas discharges rocket engines etc. will some of the atoms and molecules be ionized It is also believed that the chemistry of the earth's troposphere predominantly proceeds via reactions between neutral particles The complex system of atmospheric chemistry will be treated in one of the forthcoming volumes to this series Why then are ions considered so important that hundreds of laboratories all over the world including some of the most prestigious are involved in research programs on ions covering many different facets from biochemistry to physics One may obtain as many different answers as there are research groups busy in this field There is however one simple common feature which makes it attractive to work with ions since they carry one or more net elementary charges they can easily be guided, focused or separated by appropriate electric and magnetic fields and last but not least they can easily be detected Apart from these advantages which are welcome and appreciated by the researcher the study of molecular ions can provide

insight into very fundamental aspects of the general behavior of molecules *Ions in Solution* ,1972 **Large Ions** Tomas Baer,Cheuk-Yiu Ng,Ivan Powis,1996 Recent advances in both experimental techniques and theoretical methodologies have meant that increasingly sophisticated studies concerning the formation structures energetics and reaction dynamics of state or energy selected molecular ions can now be performed In order to better serve the ion chemistry and physics community each volume of this series is dedicated to reviewing a specific topic emphasizing new experimental and theoretical developments in the study of ions The Wiley Series in Ion Chemistry and Physics will help stimulate new research directions and point to future opportunities in the field of ion chemistry and physics This volume the sixth in the series concentrates on the area of large ions The production detection and analysis of large ions are areas which have taken on great importance in recent years in particular in the biomedical and biochemical fields The understanding of large ions presents unique and formidable challenges which are very different from those associated with the study of small ions This volume focuses on some of the fundamental chemistry and physics associated with the behavior of large ions with the contributors addressing the issues in a quantitative fashion in order to elucidate clearly some of the key recent advances which have taken place As such **Large Ions Their Vaporization Detection and Structural Analysis** provides an excellent snapshot of current research in this fascinating and important area The six chapters are written by some of the leading experts in the field and together they will be of great interest to experts and newcomers both of whom will benefit from the in depth discussion of the latest methods and results On the Relationship Between the Charge of an Ion and Its Velocity Jacob Neufeld,1957 **The Structure, Energetics and Dynamics of Organic Ions** Tomas Baer,Cheuk-Yiu Ng,Ivan Powis,1996-09-10 Recent advances in both experimental techniques and theoretical methodologies have meant that increasingly sophisticated studies concerning the formation structures energetics and reaction dynamics of state or energy selected molecular ions can now be performed In order to better serve the ion chemistry and physics community each volume of this series is dedicated to reviewing a specific topic emphasizing new experimental and theoretical developments in the study of ions The Wiley Series in Ion Chemistry and Physics will help stimulate new research directions and point to future opportunities in the field of ion chemistry and physics This volume the fifth in the series concentrates on the important area of organic ions Our understanding of these ions has been significantly improved over the last decade due to both theoretical and experimental advances and it is now routinely possible to calculate organic ion structures by ab initio molecular orbital methods with very high precision At the same time the experimental tools used to study organic ions have become much more refined This volume provides a timely overview of some of the key approaches which are currently producing such important results in the study of negative ions and contains seven chapters written by acknowledged experts in the field It will be of great interest to both experts and newcomers both of whom will benefit from the in depth discussion of the latest methods and results Survey of Multiply Charged Ions Fred Loomis Mohler,1964 **Ions in Solution (3)** Geoffrey Pass,1973

Atoms and Ions Alfred John Cork, 1969 **Ion Solvation** Y. Marcus, 1985 Chemical reactions generally take place in solution and often involve ions The behaviour of ions in solution manifested through ion solvation is therefore of prime interest in chemistry This book considers in depth the phenomenology of ion solvation and the models and interpretations that have been proposed as the physical causes for the observed phenomena It contains a thorough discussion of the statistical thermodynamic background of the solvation process from which a discussion of the actual thermodynamics is developed This in turn serves as a background to the structural and kinetic features of ion solvation **Ions in Solution**, 1974 *Nonclassical Ions* Paul D. Bartlett, 1965

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