

## Question

14. Following the redox reaction between zinc and copper ions, write balanced chemical equations for each of the following half-reactions:

Half-reaction	Reaction
1a) Zinc oxidation half-reaction	$Zn \rightarrow Zn^{2+} + 2e^-$
1b) Copper reduction half-reaction	$Cu^{2+} + 2e^- \rightarrow Cu$
1c) Zinc reduction half-reaction	$Zn^{2+} + 2e^- \rightarrow Zn$
1d) Copper oxidation half-reaction	$Cu \rightarrow Cu^{2+} + 2e^-$
1e) Zinc oxidation half-reaction	$Zn \rightarrow Zn^{2+} + 2e^-$
1f) Copper reduction half-reaction	$Cu^{2+} + 2e^- \rightarrow Cu$
1g) Zinc reduction half-reaction	$Zn^{2+} + 2e^- \rightarrow Zn$
1h) Copper oxidation half-reaction	$Cu \rightarrow Cu^{2+} + 2e^-$
1i) Zinc oxidation half-reaction	$Zn \rightarrow Zn^{2+} + 2e^-$
1j) Copper reduction half-reaction	$Cu^{2+} + 2e^- \rightarrow Cu$

15. Give three chemical reactions that demonstrate the chemical properties of zinc. Write balanced chemical equations for each reaction and label the products and reactants. (3 points)

# Ionic Bonding Chemistry If8766 Answers

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