

IONIC

DEFINITION

an electrical attraction between a positively charged ion \oplus and a negatively charged ion \ominus



more protons \oplus than electrons \ominus

atom loses electron to become \oplus charge

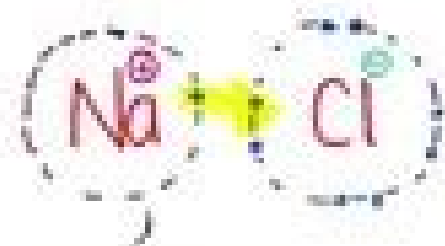


more electrons \ominus than protons \oplus

atom gains electrons to become \ominus charge

PROPERTIES

- high boiling point
- high melting point
- solid @ room temperature
- dissolve in water
- conduct electricity
- brittle



COVALENT

DEFINITION

chemical bond formed when electrons are shared.

single \longleftrightarrow share one pair of electrons

double \longleftrightarrow share two pairs of electrons

triple \longleftrightarrow shares three pairs of electrons

PROPERTIES

- do not conduct electricity
- lower melting point
- do not separate in water
- lower boiling point
- can be solid, liquid or gas at room temperature

shared electrons circle both atoms



METALLIC

DEFINITION

bond formed when many metal atoms share their pooled electrons

LIKE IONIC

give away electrons

LIKE COVALENT

metals share electrons

PROPERTIES

- high boiling point
- high melting point
- good conductors of heat & electricity
- shiny



Ionic Bonding Notes Guide

Thomas Griffiths



Ionic Bonding Notes Guide:

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