
Identify the order of the change that best completes the statement or answers the question.

- _____ 7. In many compounds, atoms of main-group elements form ions so that the number of electrons in their outermost energy levels of each ion is:

 - a. 2.
 - b. 6.
 - c. 8.
 - d. 10.
- _____ 8. An atom

 - a. is an ion with a negative charge.
 - b. attracts ions with negative charges.
 - c. results when an alkaline-earth metal loses one of its two valence-level electrons.
 - d. has more protons than electrons.
- _____ 9. When the octet rule is satisfied, the element's _____ are filled.

 - a. s and d orbitals
 - b. p and f orbitals
 - c. s and d orbitals
 - d. p and f orbitals
- _____ 10. The elements of the _____ group satisfy the octet rule without forming compounds.

 - a. halide
 - b. noble gas
 - c. alkali metal
 - d. alkaline-earth metal
- _____ 11. Which ion does not have full s and p orbitals in its outermost energy level?

 - a. It is highly reactive only with alkali metals.
 - b. It is highly reactive only with halogens.
 - c. It can be combined with most elements.
 - d. It has a stable outer shell and is unreactive.
- _____ 12. In a crystal of an ionic compound, each cation is surrounded by

 - a. molecules.
 - b. positive ions.
 - c. dipole.
 - d. anions.
- _____ 13. The lithium(I) ion and sodium(I) ion

 - a. have the same charge.
 - b. are polyatomic ions.

Ionic And Metallic Bonding Chapter 7 Worksheet Answers

R Sandford



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