ENGAGE

Why does the atomic number but not the mass number change in positron emission?

SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission

TRY IT FIRST

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

SOLUTION GUIDE

| ANALYZE THE PROBLEM | Given | Need | Connect |
|---------------------|-----------------|------------------|-----------------------|
| | Mn-49, positron | balanced nuclear | mass number, atomic |
| | emission | equation | number of new nucleus |

STEP 1 Write the incomplete nuclear equation.

$$^{49}_{29}Mm \longrightarrow 7 + ^{9}_{47}e$$

STEP 2 Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = 7 + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49$$
 (mass number of new nucleus)

STEP 3 Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = 7 + 1$$

$$25 - 1 = 7$$

$$25 - 1 = 24$$
 (atomic number of new nucleus)

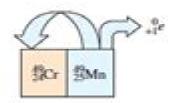
STEP 4 Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written 50r.

SIEP 5 Complete the nuclear equation.

STUDY CHECK S.4.

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

ANSWER



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H Kauffman

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Table of Contents Introduction To Mastering Chemistry Answers

- 1. Understanding the eBook Introduction To Mastering Chemistry Answers
 - The Rise of Digital Reading Introduction To Mastering Chemistry Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Introduction To Mastering Chemistry Answers
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Introduction To Mastering Chemistry Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Introduction To Mastering Chemistry Answers
 - Personalized Recommendations
 - Introduction To Mastering Chemistry Answers User Reviews and Ratings

- Introduction To Mastering Chemistry Answers and Bestseller Lists
- 5. Accessing Introduction To Mastering Chemistry Answers Free and Paid eBooks
 - Introduction To Mastering Chemistry Answers Public Domain eBooks
 - Introduction To Mastering Chemistry Answers eBook Subscription Services
 - Introduction To Mastering Chemistry Answers Budget-Friendly Options
- 6. Navigating Introduction To Mastering Chemistry Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Introduction To Mastering Chemistry Answers Compatibility with Devices
 - Introduction To Mastering Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Introduction To Mastering Chemistry Answers
 - Highlighting and Note-Taking Introduction To Mastering Chemistry Answers
 - Interactive Elements Introduction To Mastering Chemistry Answers
- 8. Staying Engaged with Introduction To Mastering Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Introduction To Mastering Chemistry Answers
- 9. Balancing eBooks and Physical Books Introduction To Mastering Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Introduction To Mastering Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Introduction To Mastering Chemistry Answers
 - Setting Reading Goals Introduction To Mastering Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Introduction To Mastering Chemistry Answers
 - Fact-Checking eBook Content of Introduction To Mastering Chemistry Answers
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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