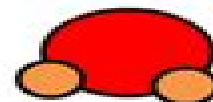


Chapter 6: Chemical Bonding Guided Notes

I. Why do Atoms Combine?

A. Chemical Formula

- The chemical formula shows:
 - The _____ in the compound.
 - The ratio of their _____

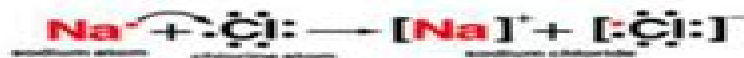


B. Chemical Bond

- Strong attractive force between _____ or ions in a molecule or compound.
- Formed by:
 - _____ electrons, (e-) by losing or gaining electrons.
 - _____ electrons (e-)

C. Stability

- Octet Rule
 - Most atoms form bonds in order to have _____ valence electrons (e-).
 - _____ outer energy level.
 - Want to be like the Noble Gases (group 18).
- Stability is the driving force behind bond formation. All atoms bond to become stable.
- Transferring e-



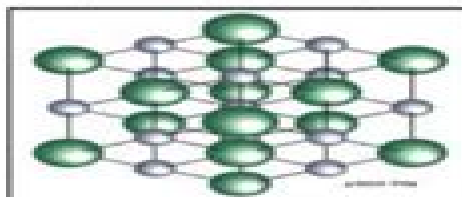
- Sharing e-



II. Kinds of Chemical Bonds

A. Ionic Bond

- Attraction between 2 _____ charged ions.
- Ions are _____ atoms.
- A _____ is a _____ charged ion. Forms when an atom loses an electron.
Formed by _____.
- A _____ is a _____ charged ion. Forms when an atom gains an electron.
Formed by _____.
- Formed by _____ electrons (e-) from a _____ to a _____.
- Ions form a 3-D crystal lattice.



Introduction To Chemical Bonding 5note Guide Key

Michael Hoelscher



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