

# Chemical Bonding

## Key Concepts

- Section 1: Introduction to Chemical Bonding**  
 What is a chemical bond and why does it form?
- Section 2: Chemical Bonding and Molecular Compounds**  
 What is a molecular formula?  
 What are the characteristics of a covalent bond?  
 How does a Lewis structure represent a molecule?
- Section 3: Ionic Bonding and Ionic Compounds**  
 What is a ionic bond?  
 How are ionic and molecular compounds different?
- Section 4: Metallic Bonding**  
 How does metallic bonding affect the properties of metals?
- Section 5: Molecular Geometry**  
 What is VSEPR theory and how is it used?  
 What other factors affect the shape of covalent molecules?

**Learning Objectives**  
 LO 1.1.1: Define and explain the terms ionic, covalent, and metallic bonds.  
 LO 1.1.2: Explain the difference between molecular and ionic compounds.  
 LO 1.1.3: Explain the difference between a chemical bond and a chemical reaction.  
 LO 1.1.4: Identify the number of bonds an atom can form and the number of valence electrons.  
 LO 1.1.5: Explain the difference between ionic and covalent bonds.  
 LO 1.1.6: Explain the difference between ionic and covalent compounds.  
 LO 1.1.7: Explain the difference between ionic and covalent compounds.  
 LO 1.1.8: Explain the difference between ionic and covalent compounds.  
 LO 1.1.9: Explain the difference between ionic and covalent compounds.  
 LO 1.1.10: Explain the difference between ionic and covalent compounds.

## Review Previous Concepts

### 1. What are valence electrons?

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### 2. What quantity does the electronegativity of an atom represent?

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