		Ю		н	
		н	ш	ш	

SS NUMBER - total amount

An atom is a really tiny particle. It has a radius of about 0.1 nanometres

(1x10-10 m) Atoms are neutral because the no. protons

= no. electrons An atom consists of:

Particle	Relative	Charg
Proton	1	+1
Neutron	1	0
Electron	0 (very	-1

You can find the

of Mumber of mestrons = atomic

#### THE NUCLEUS

- •Middle of the atom
- Contains protons and neutrons
  - Has a radius of around 1x10-14m (1/100° of atom radius)
- Positively charged due to protons
- Almost whole mass of atom concentrated in mucleus.

### THE ELECTRONS

Move around nucleus in electron shells

- •Negatively charged & tiny, but cover a lot of space
- •The volume of their orbits determines the size
- of the atoms Negatively charged
- ·Bond to other electrons to make chemical bond



### Element

A substance made up of atoms that all have the same number of protons in their nucleus (i.e. all are the same atom)

## Isotopes

Different forms of the same element, which have the same number of protons but a different number of neutrons.

## 2 common

examp -Carbon-12

\*Carbon-13

Because many elements can exist as a

number of

different isotopes... relative atomic

mass (A) is used instead of mass number

when referring

to the element

Carbon-13 Carlotte. and Property. ALL PROPERTY. # s. Combine. in Thiredness Machine anadom Market breaker 18.18 44.47

+ 111

relative alomic mass =

sum of all isotope abundances

- 10

To balance an equation, you should write out all the elements in the middle, write out how many there are on either side and add numbers to balance it

### CHEMISTRY CT

Atomic structure, separating

H2SO4 + NaOH Na/50++ H<sub>2</sub>O

BQ) Copper has two stable isotopes. Cu-63 has an abundance of 69.2% and Cu-65 has an abundance of 30.8%. Calculate the relative atomic mass of copper to 1dp.

(882-1)-(383-65)\_ 61584-2981\_43614 63.616 = 63.6

to enable

num off instage abundance + instage mumber's

The relative atomic mass is an average mass taking into account the different masses & abundances (amounts) of all the isotopes that make up the

# Compound

Substances formed from 2/more elements, where the atoms of each are held together by chemical bonds.

- Making bond involves atoms giving away, taking or sharing electrons
- A chemical bond is required to break these bonds and separate the compound
- There are 2 types of compound bonds:

### 1) IONIC BONDING

- · When a metal and a non-metal react to form compound
- The metal loses electrons to form a positive ion
- The non-metal gains electrons to form a negative ion
- The opposite charges of the ions. means they're strongly attracted

### 2) COVALENT BONDING

- A compound formed from nonmetal consists of molecules
- Each atoms shares electrons with another atom

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# **Gcse Edexcel2015c1 Chemistry**

### **Roman Wölfel**

 ${\bf Gcse} \ {\bf Edexcel 2015c1} \ {\bf Chemistry:}$ 

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1. (	Open the passenger	door and ren	nove the dash trim	on the side of the dash	. (Figure A). 2.	Open the glove box	and remove.
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