

Experiment 23

Data and Calculations: Determination of the Equilibrium Constant for a Chemical Reaction

Mixture	Volume in mL, 2.00×10^{-3} M $\text{Fe}(\text{NO}_3)_3$	Volume in mL, 2.00×10^{-3} M KSCN	Volume in mL, Water	Method I Absorbance	Method II Depth in mm		$[\text{FeSCN}^{2+}]$
					Standard	Unknown	
1	5.00	1.00	<u>4.00</u>	<u>0.084</u>	—	—	<u>0.25×10^{-4} M</u>
2	5.00	2.00	<u>3.00</u>	<u>0.218</u>	—	—	<u>0.45×10^{-4} M</u>
3	5.00	3.00	<u>2.00</u>	<u>0.317</u>	—	—	<u>0.60×10^{-4} M</u>
4	5.00	4.00	<u>1.00</u>	<u>0.447</u>	—	—	<u>0.84×10^{-4} M</u>
5	5.00	5.00	<u>0</u>	<u>0.514</u>	—	—	<u>1.00×10^{-4} M</u>

If Method II was used, $[\text{FeSCN}^{2+}]_{\text{unknown}} = \text{—} \times 10^{-4}$ M; $[\text{FeSCN}^{2+}]$ in Mixtures 1 to 5 is found by Equation 7.

Processing the Data

A. Calculation of K_c assuming the reaction:



This calculation is most easily done by following Steps 1 through 5 in the discussion. Results are to be entered in the table on the following page. If you are using Excel, set up the table as we have, and follow the directions below.

Step 1 Find the initial number of moles of Fe^{3+} and SCN^{-} in the mixtures in test tubes 1 through 5. Use Equation 3 and enter the values in the first two columns of the table.

Step 2 Enter the experimentally determined value of $[\text{FeSCN}^{2+}]$ at equilibrium for each of the mixtures in the next to last column in the table. Use Equation 3 to find the number of moles of FeSCN^{2+} in each of the mixtures, and enter the values in the fifth column of the table. Note that this is also the number of moles of Fe^{3+} and SCN^{-} that were used up in the reaction.

Experiment 23 Determination Equilibrium Constant

D Keegan



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