

1) Which of the following statements is incorrect?

- A) The degree of ionization of a weak electrolyte increases with dilution.
- B) Strong electrolytes are ionized completely at all concentrations.
- C) Addition of  $\text{NH}_4\text{Cl}$  to  $\text{NH}_4\text{OH}$  increases the ionization of the latter.
- D) Increase the temperature increases the ionization.

2) According to Lowry and Bronsted system the chloride ion ( $\text{Cl}^-$ ) in aqueous solution is a

- A) Weak base
- B) Strong base
- C) Weak acid
- D) Strong acid

3) In the following reactions,  $\text{ZnO}$  is respectively acting as a/an

i)  $\text{ZnO} + \text{Na}_2\text{O}$  gives  $\text{Na}_2\text{ZnO}_2$

ii)  $\text{ZnO} + \text{CO}_2$  gives  $\text{ZnCO}_3$

- A) Acid and acid
- B) Acid and base
- C) Base and acid
- D) Base and base

4) Given  $K_{a1}$ ,  $K_{a2}$  for  $\text{H}_2\text{S} = 1.1 \times 10^{-7} \text{M}^2$ . 1L of 0.1 M  $\text{H}_2\text{S}$  having  $[\text{H}^+] = 0.1 \text{M}$  contains  $\text{S}^{2-}$  ions equal to.

- A)  $6.62 \times 10^3$
- B)  $6.62 \times 10^4$
- C)  $6.62 \times 10^5$
- D)  $6.62 \times 10^6$

5) Equal volume of three acid solutions of pH 3, 4 and 5 are mixed in a vessel. What will be

the pH of the resulting solution?

- A) 3.33
- B) 3.5
- C) 3.67
- D) 3.83

6) The pH of a solution is 10. The concentration of  $\text{OH}^-$  ions in the solution is

- A)  $10^{-10} \text{M}$
- B)  $10^{-4} \text{M}$
- C)  $10^{-14} \text{M}$
- D)  $10^{-1} \text{M}$

7) The degree of dissociation of  $\text{H}_2\text{S}$  is 0.1. The concentration of  $\text{H}^+$  ions is

- A) 0.1
- B) 0.01
- C) 0.001
- D) 0.0001

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